

Worksheet: Periodic Trends

Name _____
Period _____

1. Which statement best describes Group 2 elements as they are considered in order from top to bottom of the Periodic Table?
(A) The number of principal energy levels increases, and the number of valence electrons increases.
(B) The number of principal energy levels increases, and the number of valence electrons remains the same.
(C) The number of principal energy levels remains the same, and the number of valence electrons increases.
(D) The number of principal energy levels remains the same, and the number of valence electrons decreases.
2. What is the total number of valence electrons in an atom of boron in the ground state?
(A) 1 (C) 3
(B) 7 (D) 5
3. What is the total number of valence electrons in an atom of xenon, Xe?
(A) 6 (C) 8
(B) 2 (D) 18
4. The elements calcium and strontium have similar chemical properties because they both have the same
(A) atomic number
(B) mass number
(C) number of valence electrons
(D) number of completely filled sublevels
5. On the Periodic Table of the Elements, all the elements within Group 16 have the same number of
(A) valence electrons (C) protons
(B) energy levels (D) neutrons
6. An element with a partially filled *d* sublevel in the ground state is classified as
(A) a halogen (C) an alkali metal
(B) a transition metal (D) an alkaline earth metal
7. Which electron configuration represents a transition element?
(A) $1s^2 2s^2 2p^2$ (C) $[\text{Ar}] 3d^5 4s^2$
(B) $[\text{Ne}] 3s^2$ (D) $[\text{Ar}] 3d^{10} 4s^2 4p^6$
8. Which element in Period 5 of the Periodic Table is a transition element?
(A) Sr (C) Ag
(B) Sb (D) Xe
9. Which of the following atoms has the largest atomic radius?
(A) Na (C) Mg
(B) K (D) Ca
10. Which noble gas has the highest first ionization energy?
(A) radon (C) neon
(B) krypton (D) helium

11. Which sequence of elements is arranged in order of decreasing atomic radii?
(A) Al, Si, P (C) Cl, Br, I
(B) Li, Na, K (D) N, C, B
12. Which list of elements from Group 2 on the Periodic Table is arranged in order of increasing atomic radius?
(A) Be, Mg, Ca (C) Ba, Ra, Sr
(B) Ca, Mg, Be (D) Sr, Ra, Ba
13. As each successive element in Group 15 of the Periodic Table is considered in order of increasing atomic number, the atomic radius
(A) decreases (C) remains the same
(B) increases
14. The strength of an atom's attraction for the electrons in a chemical bond is the atom's
(A) electronegativity (C) heat of reaction
(B) ionization energy (D) heat of formation
15. Which properties are most common in nonmetals?
(A) low ionization energy and low electronegativity
(B) low ionization energy and high electronegativity
(C) high ionization energy and low electronegativity
(D) high ionization energy and high electronegativity
16. Which Group 17 element has the least attraction for electrons?
(A) F (C) Br
(B) Cl (D) I
17. Which element in Group 16 has the greatest tendency to gain electrons?
(A) Te (C) S
(B) Se (D) O
18. The Group 17 element with the highest electronegativity is
(A) fluorine (C) bromine
(B) chlorine (D) iodine
19. As the elements of Group 1 on the Periodic Table are considered in order of increasing atomic radius, the ionization energy of each successive element generally
(A) decreases (C) remains the same
(B) increases
20. The amount of energy required to remove the outermost electron from a gaseous atom in the ground state is known as
(A) first ionization energy (C) conductivity
(B) activation energy (D) electronegativity
21. Which element is a member of the halogen family?
(A) K (C) I
(B) B (D) S

Periodic Table Trends Answer Honors

David Baud



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