

10. Circle the element in each of the following pairs with the greater (1<sup>st</sup>) ionization energy:

a) lithium or boron

c) cesium or lead

b) magnesium or strontium

d) sodium or argon

11. Arrange the following elements in order of increasing ionization energy:

a) Sr, Be, Mg

b) Bi, Cs, Ba

c) Na, Al, S

12. Why is there a very large increase between the 1<sup>st</sup> and 2<sup>nd</sup> ionization energies for the elements in group 1?

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Class period: \_\_\_\_\_



### Unit: Periodic Trends "Ionic Size and E.N. Trend" - WS #4

Directions: Please answer each fill in the blank with the best answer. ☺

1. A cation is an ion with an overall \_\_\_\_\_ charge.
2. Cations are made when an atom \_\_\_\_\_.
3. The size of a positive ion is always \_\_\_\_\_ than the size of the neutral atom of the same element.
4. As the positive charge on an ion increases, the size of the atom will \_\_\_\_\_.
5. An anion is an ion with an overall \_\_\_\_\_ charge.
6. Anions are made when at atom \_\_\_\_\_.
7. The size of a negative ion is always \_\_\_\_\_ than the size of the neutral atom of the same element.
8. As the negative charge on an ion increases, the size of the atom will \_\_\_\_\_.
9. Which element/ion has the **biggest** atomic/ionic size in each of the following pairs? Please circle your answer.
  - a) Na or Na<sup>+</sup>
  - b) S or S<sup>2-</sup>
  - c) O<sup>-1</sup> or O<sup>-2</sup>
  - d) Al or Al<sup>+3</sup>
  - e) Cu<sup>+1</sup> or Cu<sup>+2</sup>
10. \_\_\_\_\_ is the tendency for the atoms of the element to attract electrons when they are chemically combined with another element.

# Periodic Trends Unit Exam And Answers

**The Princeton Review**



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