

## Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

**Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.**

2. State the periodic law.

**The properties of the elements are periodic functions of their atomic numbers.**

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

Electron Config.	Block	Period	Group/Type	Element
[Xe] 6s <sup>2</sup> 4f <sup>14</sup> 5d <sup>10</sup> 6p <sup>6</sup>	p	6	18 / Noble Gas	Radon
[Xe] 6s <sup>1</sup>	s	6	1 / Alkali Metal	Cesium
[Ar] 4s <sup>1</sup> 3d <sup>5</sup>	d	4	6 / Trans. Metal	Chromium
[He] 2s <sup>2</sup>	s	2	2 / Alk. Earth Metal	Beryllium
[Kr] 5s <sup>1</sup> 4d <sup>10</sup>	d	5	11 / Trans. Metal	Silver
[He] 2s <sup>2</sup> 2p <sup>5</sup>	p	2	17 / Halogen	Fluorine

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- most reactive non-metals **H**
- s<sup>2</sup>p<sup>6</sup> valence configuration **NG**
- second most reactive metals **AE**
- s<sup>1</sup> valence configuration **A**
- exceptions found to e- configurations **TM**
- soft metals **A**
- d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl<sup>-</sup> ion to be larger or smaller than an Mg<sup>2+</sup> ion. Explain. **A Cl<sup>-</sup> ion is larger than a Mg<sup>2+</sup> ion, because the chloride ion has 3 energy levels, matching**

# Periodic Trends Study Guide

**James E. Brady, Fred Senese**



## Periodic Trends Study Guide:

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### **Table of Contents Periodic Trends Study Guide**

1. Understanding the eBook Periodic Trends Study Guide
  - The Rise of Digital Reading Periodic Trends Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Study Guide
  - Personalized Recommendations



- Periodic Trends Study Guide User Reviews and Ratings
- Periodic Trends Study Guide and Bestseller Lists
- 5. Accessing Periodic Trends Study Guide Free and Paid eBooks
  - Periodic Trends Study Guide Public Domain eBooks
  - Periodic Trends Study Guide eBook Subscription Services
  - Periodic Trends Study Guide Budget-Friendly Options
- 6. Navigating Periodic Trends Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Study Guide Compatibility with Devices
  - Periodic Trends Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Study Guide
  - Highlighting and Note-Taking Periodic Trends Study Guide
  - Interactive Elements Periodic Trends Study Guide
- 8. Staying Engaged with Periodic Trends Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Study Guide
- 9. Balancing eBooks and Physical Books Periodic Trends Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Study Guide
  - Setting Reading Goals Periodic Trends Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Study Guide
  - Fact-Checking eBook Content of Periodic Trends Study Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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