

Periodic Trends

ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.

a) Al B	b) S O	c) Br Cl
d) Na Al	e) O F	f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and** explain why:
C, O, Sn, Sr.

ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.

a) Ca Ga	b) Li O	c) Cl S	d) Br As	e) Ba Sr	f) O S
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GENERAL QUESTIONS

14. Which group tends to form +1 ions? _____
15. Which group tends to form +2 ions? _____
16. Which group tends to form -1 ions? _____
17. Which group tends not to form ions or react? _____

Periodic Trends Exam

Richard Bailey

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Richard Bailey.

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