

1 H Hydrogen 1.008																	2 He Helium 4.003				
3 Li Lithium 6.94	4 Be Beryllium 9.012															5 B Boron 10.81	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305															13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.06	17 Cl Chlorine 35.45	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.630	33 As Arsenic 74.922	34 Se Selenium 78.97	35 Br Bromine 79.904	36 Kr Krypton 83.798				
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.94	43 Tc Technetium [97]	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.757	52 Te Tellurium 127.6	53 I Iodine 126.905	54 Xe Xenon 131.29				
55 Cs Cesium 132.905	56 Ba Barium 137.327	★ 57 - 71	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	78 Ir Iridium 192.222	79 Pt Platinum 195.084	80 Au Gold 196.967	81 Hg Mercury 200.592	82 Tl Thallium 204.38	83 Pb Lead 207.2	84 Bi Bismuth 208.980	85 Po Polonium [209]	86 At Astatine [210]	88 Rn Radon [222]				
87 Fr Francium [223]	88 Ra Radium [226]	★★ 89 - 103	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [277]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [271]	111 Rg Roentgenium [272]	112 Cn Copernicium [285]	113 Nh Nihonium [284]	114 Fl Flerovium [289]	115 Mc Moscovium [288]	116 Lv Livermorium [293]	117 Ts Tennessine [294]	118 Og Oganesson [294]				

Metals

Nonmetals

Metaloids

Reactor

Accelerator

Atomic number

Symbol

Name

Average atomic mass

6


C

Carbon

12.011

Atomic number — 6
 Symbol — C
 Name — Carbon
 Average atomic mass — 12.011

Metals 
 Nonmetals 
 Metalloids 

 Reactor 
 Accelerator 

*Lanthanide series

57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.242	61 Pm Promethium [145]	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.502	67 Ho Holmium 164.930	68 Er Erbium 167.256	69 Tm Thulium 168.934	70 Yb Ytterbium 173.046	71 Lu Lutetium 174.967
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**Actinide series

89 Ac Actinium [227]	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium [237]	94 Pu Plutonium [244]	95 Am Americium [243]	96 Cm Curium [247]	97 Bk Berkelium [247]	98 Cf Californium [251]	99 Es Einsteinium [252]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]	103 Lr Lawrencium [262]
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Periodic Table Basics Key

R Barnett



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Analytical Chemistry Bryan M. Ham,Aihui MaHam,2024-02-28 ANALYTICAL CHEMISTRY Detailed reference covering all aspects of working in laboratories including safety fundamentals of analytical techniques lab instrumentation and more A comprehensive study of analytical chemistry as it pertains to the laboratory analyst and chemist Analytical Chemistry begins with an introduction to the laboratory environment including safety glassware common apparatuses and lab basics and continues on to guide readers through the fundamentals of analytical techniques such as spectroscopy and chromatography and introduce examples of laboratory programs such as Laboratory Information Management Systems LIMS This newly updated and revised Second Edition of Analytical Chemistry offers expanded chapters with new figures and the latest developments in the field Included alongside this new edition is an updated companion teaching reference and toolkit program called ChemTech Conveniently available via either app or browser the ChemTech program contains exercises that highlight and review topics covered in the book and features useful calculators and programs including solution makers graphing tools and more To aid in reader comprehension the program also includes an interactive periodic table and chapter summaries Written by two highly qualified authors with significant experience in both practice and academia Analytical Chemistry covers sample topics such as Basic mathematics in the laboratory including different units the metric system significant figures scientific calculators and ChemTech conversion tools Analytical data treatment including errors in the laboratory precision versus accuracy normal distribution curves and determining errors in methodology Plotting and graphing including graph construction curve fitting graphs of specific equations least squares method and computer generated curves Ultraviolet visible UV Vis spectroscopy including wave and particle theory of light light absorption transitions the color wheel and pigments With complete coverage of the practical aspects of analytical chemistry Analytical Chemistry prepares students for a rewarding career as a chemist or a laboratory technician Thanks to ChemTech integration the book is also a useful and accessible reference for the established chemist or technician already working in the laboratory

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learn how to get the most from this guide Chapter 1 Foundations of Science Core principles the scientific method and the importance of hypotheses and theories Chapter 2 History of Science Major eras and groundbreaking discoveries that define modern science Chapter 3 Physics Matter energy mechanics thermodynamics and electromagnetism Chapter 4 Chemistry

Matter s properties chemical reactions bonding and the periodic table Chapter 5 Biology Genetics evolution ecosystems and cellular biology Chapter 6 Earth Science Geology meteorology oceanography and environmental conservation Chapter 7

Astronomy Our solar system galaxies black holes and cosmology Chapter 8 Human Body Medical Science Human anatomy disease prevention and medical advances Chapter 9 Environmental Science Human impact on the planet sustainability and climate change Chapter 10 Mathematics in Science Math s role in research from probability to calculus Chapter 11 Social

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Table of Contents Periodic Table Basics Key

1. Understanding the eBook Periodic Table Basics Key
 - The Rise of Digital Reading Periodic Table Basics Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Basics Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Basics Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Basics Key

- Personalized Recommendations
- Periodic Table Basics Key User Reviews and Ratings
- Periodic Table Basics Key and Bestseller Lists
- 5. Accessing Periodic Table Basics Key Free and Paid eBooks
 - Periodic Table Basics Key Public Domain eBooks
 - Periodic Table Basics Key eBook Subscription Services
 - Periodic Table Basics Key Budget-Friendly Options
- 6. Navigating Periodic Table Basics Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Basics Key Compatibility with Devices
 - Periodic Table Basics Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Basics Key
 - Highlighting and Note-Taking Periodic Table Basics Key
 - Interactive Elements Periodic Table Basics Key
- 8. Staying Engaged with Periodic Table Basics Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Basics Key
- 9. Balancing eBooks and Physical Books Periodic Table Basics Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Basics Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Basics Key
 - Setting Reading Goals Periodic Table Basics Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Basics Key

-
- Fact-Checking eBook Content of Periodic Table Basics Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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