

Quiz & Worksheet - Atomic and Ionic Radii on the Periodic Table

1. As you move from left to right on the periodic table, why does the atomic radius decrease?

- As more protons are added, the more tightly the surrounding electrons are drawn inwards.
- As more neutrons are added, the more tightly the surrounding electrons are drawn inwards.
- As more electrons are added, the more tightly they draw themselves towards the nucleus.
- As the mass of the nucleus rises, the more tightly the surrounding electrons are drawn inwards.

2. As you move from top to bottom on the periodic table, what happens to the atomic radius?

- in. It decreases.
- It decreases, then increases.
- tt nemains constant.
- It increases, then decreases.
- It increases.

3. Which of these elements has the largest atomic radius: argon, phosphorus, aluminum, sodium or magnesium?



- in: Argon (Art)
- Phosphorus (P)
- Seedium (Nati
- Altuminum (Alb
- Magnesium (Mg)

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Periodic Elements Need To Knows Answer Key:

Unveiling the Energy of Verbal Artistry: An Psychological Sojourn through **Periodic Elements Need To Knows Answer Key**

In a global inundated with screens and the cacophony of instant interaction, the profound energy and psychological resonance of verbal artistry often disappear into obscurity, eclipsed by the continuous onslaught of sound and distractions. Yet, set within the musical pages of **Periodic Elements Need To Knows Answer Key**, a captivating function of literary splendor that pulses with raw feelings, lies an unforgettable journey waiting to be embarked upon. Published by a virtuoso wordsmith, this mesmerizing opus books viewers on a mental odyssey, softly revealing the latent potential and profound impact embedded within the complex web of language. Within the heart-wrenching expanse of this evocative analysis, we shall embark upon an introspective exploration of the book is key styles, dissect their captivating writing style, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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