

Percent Yield

Zinc reacts with iodine in a synthesis reaction:



- Determine the theoretical yield if 1.912 mol of zinc is used.
- Determine the percent yield if 515.6 g of product is recovered.

$$\frac{\text{AY}}{\text{TY}} \times 100 = \% \text{ yield} \quad \text{A.Y.}$$

$$\frac{515.6 \text{ g}}{610.29 \text{ g}} \times 100 = 84.48\% \text{ yield ZnI}_2$$

$$\text{T.Y.} = 610.29 \text{ g ZnI}_2$$

Percent Yield Holt Chemfile Answers

Hongru Du

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

Holt Chemistry Ralph Thomas Myers,2004 Holt Chemistry ,2003 **Holt ChemFile Lab Program** ,2005

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