REDUCTION-OXIDATION REACTIONS

Q1 - Identifying Oxidation and Reduction

Assign exidation numbers to each element in the following equations then identify:

- oxidation and reduction
- oxidising and reducing agents

Los or agent

CO Frest Ogent

- c) H,S . (NO.) > S . NO . H,O
- d) Cd + NIO, + H,O Cd(OH), + NI(OH),

Q2 - Balancing by Half Equations

Balance each of the following redox reactions using the half equation method.

- a) (Cr2O7)2 + CT -> Cr2 + Cl2 (acidic) (Cr2O2)2 . SCI + 14H1 -= 2Cr2 , 3Cl. + 7H50
- b) Cu + (NO₁) → Cu² + NO₂ (ecidic) Cu + 2(NO₂) + 4H → Cu²⁺ + 2NO + ZH₂O
- 0) (NO) ALAS TSH) AK(OH) (Duste) (NO2) + 2AI + 5H2O + (OH) => 2AI(OH)4- + NH3
- d) (MnO4) + CH2OH -> Mn2 + HCO2H (acidic) 4 (RHO4) + 5CHCOH + 12H + 54KH+ 140
- e) (NO2) + (Gr2O2)2 + Gr2 + (NO2) (acidic) 3(UO2) + (G2O2)2 + 8H + 2(F2 + 3(NO2)2 + 4H0
- f) cr(OH); + (ClO) -> (CrO₄)2 + Cl₂ (basic) = (r(OH); + 6 (ClO) -> 2 (CrO₄)2 + 8CL +24,0 +2(OH)
- g) Iz + (CIO) + (IO) + CI (acidic) Iz + 5(CIO) + +10- 2(D) +5CIT +24+
- h) AsiO1 + (NO1) + H1ASO4 + N2O1 (ocidic) AsiO1 + 2(NO) + 2H2O+2H+ + 2H2HO2 + N2O1
- i) H2Oz + ClOz + (ClOz) + Oz (basic) H6Oz + 2ClOz + 2(OA) 2(ClOz) + Oz + 2H2O
- 1) (Cr2Or)2 . I -> Cr2 . (IO2) (scidic) (CrO2)2 . I + 8H1 = 2Cr2 + (ID2) . 9H30
- k) (MnO1) + Cl + Mn2 + Cl2 (acidic) 2 (HnQ1) + 10Cl , 16Ht = 2H1= 1 5Cl2+ 8+0
- 1) (MnO) + Br + MnO2 + (BrO2) (basic) 2(MnO2) + Br + HO + 2MrO2 + (BrO2) +2(OH)
- m) As + (ClO2) H2ASO4 + HClO (acidic) 4 A . 5(00) 6 HO. 5H 4 HAQ + 5HCIO
- n) (Cr2O2)2 + CH2OH + Cr2 + HCO2H (ocidic) 4(Cr2O2)2 , 9(H2OH -20H -2 BCC2 , 9 HO2H -19 H2O
- o) Pb(OH)+2 + (ClO) → PbO; + CT (basic) Pb (OH), + (ClO) PbQ + CI+H₂O , 2 (OH) -
- p) Mn3 + NaBiO3 -> Bi3 + (MnO4) + Na (acidic) 2 Hn2 + 5 NoB; O3 + (Lift 5 E12 + 2 (HnO4)
- q) $H_2O_2 + Cl_2O_2 \rightarrow (ClO_2)^2 + O_2$ (basic)

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Table of Contents Oxidation Reduction Concept Review Answers

- 1. Understanding the eBook Oxidation Reduction Concept Review Answers
 - The Rise of Digital Reading Oxidation Reduction Concept Review Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Oxidation Reduction Concept Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Oxidation Reduction Concept Review Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Oxidation Reduction Concept Review Answers
 - Personalized Recommendations
 - Oxidation Reduction Concept Review Answers User Reviews and Ratings
 - Oxidation Reduction Concept Review Answers and Bestseller Lists

- 5. Accessing Oxidation Reduction Concept Review Answers Free and Paid eBooks
 - Oxidation Reduction Concept Review Answers Public Domain eBooks
 - Oxidation Reduction Concept Review Answers eBook Subscription Services
 - Oxidation Reduction Concept Review Answers Budget-Friendly Options
- 6. Navigating Oxidation Reduction Concept Review Answers eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Oxidation Reduction Concept Review Answers Compatibility with Devices
 - Oxidation Reduction Concept Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Oxidation Reduction Concept Review Answers
 - Highlighting and Note-Taking Oxidation Reduction Concept Review Answers
 - Interactive Elements Oxidation Reduction Concept Review Answers
- 8. Staying Engaged with Oxidation Reduction Concept Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Oxidation Reduction Concept Review Answers
- 9. Balancing eBooks and Physical Books Oxidation Reduction Concept Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Oxidation Reduction Concept Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Oxidation Reduction Concept Review Answers
 - Setting Reading Goals Oxidation Reduction Concept Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Oxidation Reduction Concept Review Answers
 - Fact-Checking eBook Content of Oxidation Reduction Concept Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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