

#### 4) Oxidation-Reduction Reactions

**Oxidation** - is a loss of electrons  
 - is an increase in oxidation number

**Reduction** - is a gain of electrons  
 - is a decrease in oxidation number

**Oxidizing Agent** - brings about the oxidation of another species  
*↑ gets reduced*

**Reducing Agent** - brings about the reduction of another species  
*↑ causes something else to gain electrons by losing electrons itself (Is Oxidized)*  
*↑ gets oxidized*

★ OIL RIG (Oxidation Is Loss; Reduction Is Gain)  
 LEO-GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal  $\xrightarrow{\text{ox}}$  metal ion  
 $2n \rightarrow 2n^{2+}$
2. Metal  $\xrightarrow{\text{ox}}$  metallic ion  $\xrightarrow{\text{ox}}$  metallic ion  
 $\text{Cu} \rightarrow \text{Cu}^+ \rightarrow \text{Cu}^{2+}$
3. Halide ion  $\xrightarrow{\text{ox}}$  halogen  $\xrightarrow{\text{ox}}$  halite ion  $\xrightarrow{\text{ox}}$  halate ion (similar for  $\text{NO}_2^- \xrightarrow{\text{ox}} \text{NO}_3^-$ )  
 $\text{Cl}^- \rightarrow \text{Cl}_2 \rightarrow \text{ClO}^- \rightarrow \text{ClO}_2^- \rightarrow \text{ClO}_3^- \rightarrow \text{ClO}_4^-$

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex.  $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$  (acidified solution)  
 $\text{MnO}_4^- \rightarrow \text{MnO}_2$  (basic solution)  
 $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$  (acidified solution)

===Identify the substances being oxidized and reduced in each of the following.===



MEMORIZE!

# Oxidation Reduction Answers

**C Cleary**



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