lame: _____

Oxidation Numbers Worksheet

Directions: Use the Rules for Assigning Oxidation Numbers to determine the oxidation number assigned to each element in each of the given chemical formulas.

Rules for Assigning Oxidation Numbers	Example:
The oxidation number of any uncombined element is 0.	Ne, O₂ > 0
2. The oxidation number of a monatomic ion equals the charge on the ion.	Fe*3 - +3
3. The oxidation number of halogens in a compound is mostly -1.	F-11
4. Oxygen has an oxidation number of -2 unless it's a peroxide is -1	H ₂ O ⁻² = -2 Na2O2 ⁻² = -1
 The oxidation number of a metal is +1 in Group 1 and +2 in Group 2. 	K *1 = +1
 Hydrogen works with *1 with nonmetals and -1 with metals. 	H*1Cl* +1 Na*1H*1 = -1
The sum of the oxidation numbers of all atoms in a neutral compound is 0.	
8. The sum of the oxidation numbers of all atoms in a polyatomic ion = the char	ge of the ion.

1. Give oxidation numbers for the underlined atoms in these molecules:

a. Cs ₂	о с	s:	0:		i. N ₂	N:	
b. N ₂ C	D ₃ N	E .	O:		j. Kr	Kr:	
c. Na	SiO ₄ N	la:	Si:	O:	k. H₂O	H:	O:
d. K ₂ C	ir ₂ O ₇ K		Cr:	0:	L. FeO	Fe:	O:
e. H ₂ C) ₂ Н	t .	O: (This	is a peroxide)	m. CaS	Ca:	S:
f. Al(OH)3 A	Д:	0	H:	n. H ₂	H:	
g. HPI	О ₃ Н		P:	O:	o. He	He:	
h. H ₂ S	ieO ₃ H	6	Se:	O:	p. H ₂ SO ₄	S:	H:

2. Give the oxidation numbers for the following ions

a. Cu ⁻¹	Cu:		b. Co ²⁺	Co:		c. Cl ⁻¹	CL	
d. IO ₂ -1	16	0:	e. SbFe -1	Sb:	F:	f. OH 1	O:	Н:

Oxidation Number Problems With Answers

Christian G. Meyer

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