

Unit 7: Reduction, Oxidation and Electrochemistry

Chapter 17: Electrochemistry

4.9: Oxidation-Reduction Reactions

Reduction-Oxidation Reactions (Redox Rxn): - chemical reactions where there is a transfer of electron(s).

Oxidation States (Oxidation Number): - a number that is arbitrary assigned to an atom in an element, molecule, and polyatomic ions to account for the number of electrons in *redox* reaction.

Rules for Assigning Oxidation Numbers:

1. **Natural Elements (include diatomic and polyatomics) have an Oxidation Number of 0.**

Examples: $\text{Na}_{(s)}$, $\text{O}_{2(g)}$, $\text{O}_{3(g)}$, $\text{H}_{2(g)}$, $\text{F}_{2(g)}$, $\text{P}_{4(s)}$, and $\text{Hg}_{(l)}$ all have an oxidation number of 0.

2. **Single Atomic Ions have an Oxidation Number Equals to its Charge.**

Example: K^+ has an oxidation number of +1.

3. **Oxygen in Binary Compound and Polyatomic Ions has an Oxidation Number of -2.**

Examples: CO_2 , CO , SO_3 , SO_2 , CO_3^{2-} , SO_4^{2-} and H_2O all have an oxidation number of -2 for oxygen.

An exception occurs in peroxides (compound containing O_2^{2-}) where the oxygen has an oxidation number of -1.

4. **Hydrogen in Binary Compounds and Polyatomic Ions has an Oxidation Number of +1.**

Examples: H_2O , HCl , HBr , CH_4 , NH_4^+ and H_2S all has an oxidation number of +1 for hydrogen.

5. **In Binary Compounds, the Atom that has the most attraction to electrons (HIGHER Electronegativity) is assigned the NEGATIVE Oxidation Number. This Negative Oxidation Number is the SAME as its Ionic Charge.**

Examples: HF : F is more electronegative. Thus, F has an oxidation number of -1.

NH_3 : N is more electronegative. Thus, N has an oxidation number of -3.

H_2S : S is more electronegative. Thus, S has an oxidation number of -2.

6. **The Sum of all Oxidation Numbers is 0 for all Electrically Neutral Compounds.**

Examples: CO : Oxidation numbers: O = -2; C = +2 $[(+2) + (-2) = 0]$

SO_2 : Oxidation numbers: O = -2; S = +4 $[(+4) + 2(-2) = 0]$

7. **The Sum of all Oxidation Numbers is Equal to the Overall Charge for all Polyatomic Ions.**

Examples: CO_3^{2-} : Oxidation numbers: O = -2; C = +4 $[(+4) + 3(-2) = -2]$

SO_4^{2-} : Oxidation numbers: O = -2; S = +6 $[(+6) + 4(-2) = -2]$

Note: Same Atom from different chemical species might NOT have the Same Oxidation Number.

We write the charges of ionic species as n^+ or n^- . Oxidation numbers are written as $+n$ or $-n$.

Oxidation Reduction Study Guide Answers

Léo Pomar



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