

- (b) A solution is prepared by the addition of 10.0 mL of 0.100 M CH_3COOH and 10.0 mL of 0.100 M NaOH .

- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

The total volume is 20.0 mL. The concentration of CH_3COOH is 0.050 M.

In this buffer, we should consider both of the species involved that would be involved in including appropriate species and correct calculation.

For CH_3COOH and CH_3COO^- and H^+ and OH^- .

$$[\text{CH}_3\text{COOH}] = 10^{-2} \times 0.100 = 1.0 \times 10^{-3} \text{ mol/L}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{1.0 \times 10^{-3}} = \frac{1.75 \times 10^{-5}}{1.0 \times 10^{-3}} = 1.75 \times 10^{-2} \text{ mol/L}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.75 \times 10^{-2}) = 1.75$$

$$\text{pH} = 1.75$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$
Initial: 0.100 M, 0.000 M, 0.000 M
Change: -0.050 M, +0.050 M, +0.050 M
Equilibrium: 0.050 M, 0.050 M, 0.050 M
pH = 1.75

- (b) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M.

In this buffer, we should consider both of the species involved that would be involved in including appropriate species and correct calculation.



- (c) The solution is a buffer solution of CH_3COOH and CH_3COO^- . The concentration of CH_3COOH is 0.050 M.

In this buffer, we should consider both of the species involved that would be involved in including appropriate species and correct calculation.

- (d) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

The concentration of CH_3COOH is 0.050 M.

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{0.050} = 1.75 \times 10^{-5}$$

$$\text{pH} = -\log[\text{H}^+] = 4.75$$

$$\text{pH} = 4.75$$

- (e) The solution is a buffer solution of CH_3COOH and CH_3COO^- .

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{0.050} = 1.75 \times 10^{-5}$$

The concentration of CH_3COOH is 0.050 M.

The concentration of CH_3COO^- is 0.050 M.

$$\text{pH} = 4.75$$

$$\text{pH} = 4.75$$

$$\text{pH} = 4.75$$

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