

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equilibrium is at 1 bar and $T_f = 1.00^\circ C$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180.16 \text{ g}}}{0.100 \text{ kg}} = 5.55 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 5.55}{1000} = 1.04^\circ C \text{ freezing}$$

$\Delta T_f = 1.86^\circ C \times 5.55 \text{ mol/kg} = 1.04^\circ C$
 $T_f = 1.00^\circ C - 1.04^\circ C = -0.04^\circ C$
 $\Delta T_b = 0.51^\circ C \times 5.55 \text{ mol/kg} = 2.83^\circ C$
 $T_b = 100^\circ C + 2.83^\circ C = 102.83^\circ C$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equilibrium is at 1 bar and $T_f = 1.00^\circ C$. $K_f = 1.86^\circ C$, $K_b = 0.51^\circ C$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60.06 \text{ g}}}{0.100 \text{ kg}} = 16.65 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 16.65}{1000} = 3.09^\circ C \text{ freezing}$$

$\Delta T_f = 1.86^\circ C \times 16.65 \text{ mol/kg} = 3.09^\circ C$
 $T_f = 1.00^\circ C - 3.09^\circ C = -2.09^\circ C$
 $\Delta T_b = 0.51^\circ C \times 16.65 \text{ mol/kg} = 8.49^\circ C$
 $T_b = 100^\circ C + 8.49^\circ C = 108.49^\circ C$

3. A 100 g mass of water at $100^\circ C$ is added to 100 g of water at $20^\circ C$. Assuming that the concentration of the solution is homogeneous, $100^\circ C$ is the boiling-point of the solution (1 bar). What is the boiling-point of the solution (1 bar) if the concentration of the solution is homogeneous?

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180.16 \text{ g}}}{0.100 \text{ kg}} = 5.55 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = \frac{1.86^\circ C \times 5.55}{1000} = 1.04^\circ C$$

$$T_f = 1.00^\circ C - 1.04^\circ C = -0.04^\circ C$$

$$T_b = 102.83^\circ C$$

$$\Delta T_b = \frac{K_b \times \text{molality}}{1000} = \frac{0.51^\circ C \times 5.55}{1000} = 2.83^\circ C$$

$$T_b = 100^\circ C + 2.83^\circ C = 102.83^\circ C$$

$$\Delta T_b = 0.51^\circ C \times 5.55 \text{ mol/kg} = 2.83^\circ C$$

$$\Delta T_b = \frac{K_b \times \text{molality}}{1000} = \frac{0.51^\circ C \times 5.55}{1000} = 2.83^\circ C$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is composed. The freezing-point of this solution is $1.0^\circ C$. The normal freezing-point of pure water is $0^\circ C$ and the freezing-point of pure water is $100^\circ C$. Assuming the total mass of the compound is 100 g, what is the boiling-point of the solution (1 bar)?

$$\Delta T_f = \frac{K_f \times \text{molality}}{1000} = 1.86^\circ C \times \text{molality} = 1.0^\circ C \Rightarrow \text{molality} = 0.538 \text{ mol/kg}$$

$$\Delta T_b = \frac{K_b \times \text{molality}}{1000} = 0.51^\circ C \times 0.538 \text{ mol/kg} = 0.274^\circ C$$

$$T_b = 100^\circ C + 0.274^\circ C = 100.274^\circ C$$

$$T_b = 100.274^\circ C$$

$$\Delta T_b = \frac{K_b \times \text{molality}}{1000} = 0.51^\circ C \times 0.538 \text{ mol/kg} = 0.274^\circ C$$

$$T_b = 100^\circ C + 0.274^\circ C = 100.274^\circ C$$

$$\Delta T_b = 0.51^\circ C \times 0.538 \text{ mol/kg} = 0.274^\circ C$$

$$\Delta T_b = \frac{K_b \times \text{molality}}{1000} = \frac{0.51^\circ C \times 0.538}{1000} = 0.274^\circ C$$

$$T_b = 100^\circ C + 0.274^\circ C = 100.274^\circ C$$

Rosengarten Colligative Properties Homework Answers

Hussin A.Rothana



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