

Name: _____

Oxidation Numbers Worksheet

Directions: Use the *Rules for Assigning Oxidation Numbers* to determine the oxidation number assigned to each element in each of the given chemical formulas.

Rules for Assigning Oxidation Numbers	Example:
1. The oxidation number of any uncombined element is 0 .	Ne, O ₂ = 0
2. The oxidation number of a monatomic ion equals the charge on the ion.	Fe ⁺³ = +3
3. The oxidation number of halogens in a compound is mostly -1 .	F ⁻¹ = -1
4. Oxygen has an oxidation number of -2 unless it's a peroxide is -1	H ₂ O ⁻² = -2 Na ₂ O ₂ ⁻² = -1
5. The oxidation number of a metal is +1 in Group 1 and +2 in Group 2.	K ⁺¹ = +1
6. Hydrogen works with +1 with nonmetals and -1 with metals.	H ⁺¹ Cl ⁻¹ = +1 Na ⁺¹ H ⁻¹ = -1
7. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	
8. The sum of the oxidation numbers of all atoms in a polyatomic ion = the charge of the ion.	

1. Give oxidation numbers for the underlined atoms in these molecules:

- | | | | | | | |
|--|-----|-----|-------------------------|---------------------------------------|-----|----|
| a. <u>Cs₂O</u> | Cs: | O: | i. <u>N₂</u> | N: | | |
| b. <u>N₂O₃</u> | N: | O: | j. <u>Kr</u> | Kr: | | |
| c. <u>Na₄SiO₄</u> | Na: | Si: | O: | k. <u>H₂O</u> | H: | O: |
| d. <u>K₂Cr₂O₇</u> | K: | Cr: | O: | l. <u>FeO</u> | Fe: | O: |
| e. <u>H₂O₂</u> | H: | O: | (This is a peroxide) | m. <u>CaS</u> | Ca: | S: |
| f. <u>Al(OH)₃</u> | Al: | O: | H: | n. <u>H₂</u> | H: | |
| g. <u>HPO₃</u> | H: | P: | O: | o. <u>He</u> | He: | |
| h. <u>H₂SeO₃</u> | H: | Se: | O: | p. <u>H₂SO₄</u> | S: | H: |

2. Give the oxidation numbers for the following ions

- | | | | | | | |
|---------------------------------------|-----|---------------------------|--|---------------------------|-----|----|
| a. <u>Cu⁺¹</u> | Cu: | b. <u>Co²⁺</u> | Co: | c. <u>Cl⁻¹</u> | Cl: | |
| d. <u>IO₂⁻¹</u> | I: | O: | e. <u>SbF₆⁻¹</u> | Sb: | F: | |
| | | | | f. <u>OH⁻¹</u> | O: | H: |

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