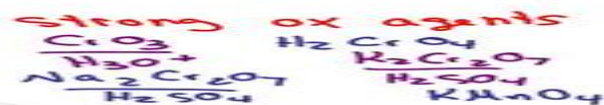


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CCCCO $\xrightarrow{\text{PCC}}$ CCCC=O
CCCCO $\xrightarrow{\text{Strong Ox agents}}$ CCCC(=O)O



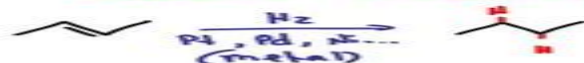
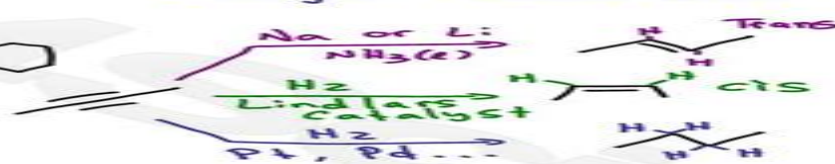
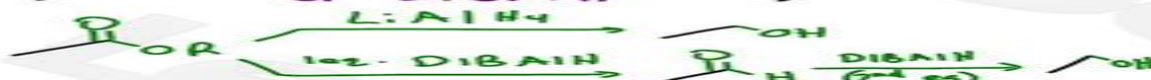
CC=CC
 $\xrightarrow[\text{KOH, } \text{O}_4]{\text{OsO}_4}$
CC(O)CC(O)



Chemical reaction showing the oxidation of 4-ethyl-2-methylstyrene using KMnO_4 to form 4-carboxy-2-methylbenzoic acid.

only 1° & 2°
get oxidized
3° = N/A

Alkene Reduction

c1ccccc1 $\xrightarrow[\text{200°C, NH}_3]{\text{Li, or Na}}$ c1ccccc1CC(=O)O>[LiAlH4]>CCO

$\text{Li}^+ + \text{H}-\ddot{\text{N}}(\text{H})-\text{H} \rightarrow \text{NaOH} + \text{H}-\ddot{\text{N}}(\text{H})-\text{H}$
 $\text{Na}^+ + \text{H}-\ddot{\text{N}}(\text{H})-\text{H} \rightarrow \text{NaOH} + \text{H}-\ddot{\text{N}}(\text{H})-\text{H}$



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Oxidation Reduction Reactions Study Guide

Anthony S. Fauci



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