

## Nobel - OCR: Chemistry Unit 2 – Chains, Energy and Resources

(a) Interpret and use the terms:

(i) Empirical formula = simplest whole number ratio of atoms of each element present in a compound

(ii) Molecular formula = actual number of atoms of each element in a molecule,

(iii) General formula = simplest algebraic formula of a member of a homologous series

Alkane	$C_nH_{2n+2}$
Alkene	$C_nH_{2n}$
Alcohol	$C_nH_{2n+2}OH$
Cycloalkane	$C_nH_{2n}$
Cycloalkene	$C_nH_{2n-2}$
Haloalkane	$C_nH_{2n+2}X$

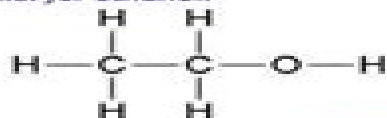
(iv) Structural formula = minimal detail that shows the arrangement of atoms in a molecule

E.g. Butane:

$CH_3CH_2CH_2CH_3$  or  $CH_3(CH_2)_2CH_3$ .

(v) Displayed formula = relative positioning of atoms and the bonds between them,

i.e. for ethanol:



(vi) Skeletal formula = simplified organic formula, shown by removing hydrogen atoms from alkyl chains, leaving just a carbon skeleton and associated functional group

ie for butan-1-ol:



(b) Interpret, and use, the terms:

(i) homologous series = a series of organic compounds having the same functional group but with each successive member differing by  $CH_2$ .

(ii) functional group = a group of atoms responsible for the characteristic reactions of a compound;

(c) Use the general formula of a homologous series to predict the formula of any member of the series;

E.g – Alkane =  $C_nH_{2n+2}$

If  $C = 9$ ,  $H = (9 \times 2) + 2$

Therefore =  $C_9H_{20}$  = Nonane

(d) State the names of the first ten members of the alkane homologous series;

No of Carbons	Alkyl Group
1	Methyl
2	Ethyl
3	Propyl
4	Butyl
5	Pentyl
6	Hexyl
7	Heptyl
8	Octyl
9	Nonyl
10	Decyl

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**O García**



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## **Table of Contents Ocr Chemistry Unit F32june 2014 Past Paper**

1. Understanding the eBook Ocr Chemistry Unit F32june 2014 Past Paper
  - The Rise of Digital Reading Ocr Chemistry Unit F32june 2014 Past Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry Unit F32june 2014 Past Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Chemistry Unit F32june 2014 Past Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry Unit F32june 2014 Past Paper
  - Personalized Recommendations
  - Ocr Chemistry Unit F32june 2014 Past Paper User Reviews and Ratings
  - Ocr Chemistry Unit F32june 2014 Past Paper and Bestseller Lists
5. Accessing Ocr Chemistry Unit F32june 2014 Past Paper Free and Paid eBooks
  - Ocr Chemistry Unit F32june 2014 Past Paper Public Domain eBooks
  - Ocr Chemistry Unit F32june 2014 Past Paper eBook Subscription Services
  - Ocr Chemistry Unit F32june 2014 Past Paper Budget-Friendly Options

6. Navigating Ocr Chemistry Unit F32june 2014 Past Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Chemistry Unit F32june 2014 Past Paper Compatibility with Devices
  - Ocr Chemistry Unit F32june 2014 Past Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Chemistry Unit F32june 2014 Past Paper
  - Highlighting and Note-Taking Ocr Chemistry Unit F32june 2014 Past Paper
  - Interactive Elements Ocr Chemistry Unit F32june 2014 Past Paper
8. Staying Engaged with Ocr Chemistry Unit F32june 2014 Past Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry Unit F32june 2014 Past Paper
9. Balancing eBooks and Physical Books Ocr Chemistry Unit F32june 2014 Past Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry Unit F32june 2014 Past Paper
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry Unit F32june 2014 Past Paper
  - Setting Reading Goals Ocr Chemistry Unit F32june 2014 Past Paper
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry Unit F32june 2014 Past Paper
  - Fact-Checking eBook Content of Ocr Chemistry Unit F32june 2014 Past Paper
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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