

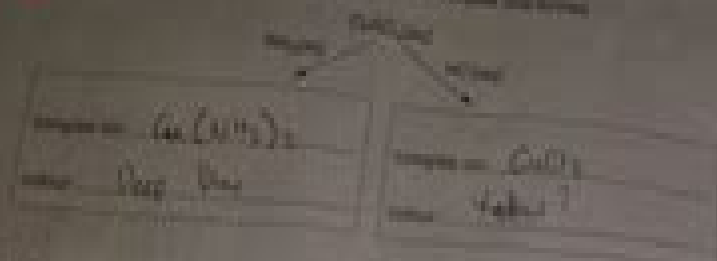
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and solvent.



(b) Cu^{2+} ions form a complex ion B with two ethylenediamine ligands and two water molecules. The ethylenediamine ion is a bidentate ligand.

The skeletal formula of the ethylenediamine ion is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand with two donor atoms that can form two coordinate bonds with a central metal ion.

(d) The complex ion A is shown in Fig. 1.2 below.

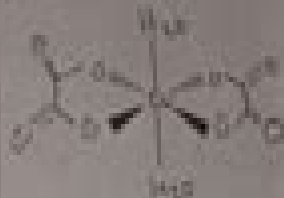
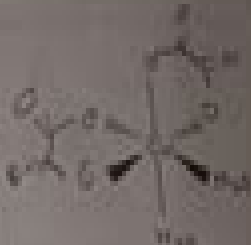
The structure of one of the ethylenediamine ligands is shown below. The charge has been given.



Complex A

Complete the 3D diagram of the other two components of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Diagram														
Type	<table><tr><td>cis</td><td></td></tr><tr><td>trans</td><td>☒</td></tr><tr><td>optical</td><td></td></tr></table>	cis		trans	☒	optical		<table><tr><td>cis</td><td>☒</td></tr><tr><td>trans</td><td></td></tr><tr><td>optical</td><td>☒</td></tr></table>	cis	☒	trans		optical	☒
cis														
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optical														
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trans														
optical	☒													

(e) What is the empirical formula, including the charge, of the complex ion A?

$[\text{Cu}(\text{NH}_2)_2(\text{NH}_3)_2]^{2+}$

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Thomas Griffiths



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