

## CHEMISTRY Objective Final Exam Review 10-11

*There are 50 questions on the test. The chapters are listed below. Each of the letters represents a concept/idea that you should know or look up.*

1. CH 6: Periodic trends, where do they increase & decrease
  - a. Atomic radius & Ionic radius
  - b. Electronegativity
  - c. Electron affinity
  - d. Ionization energy
  - e. Reactivity of metals and of non-metals
2. CH 7: Look up properties of elements per group and the names of each group.
  - a. Group 1 very reactive in water; group 7 most reactive non-metals etc.
  - b. Fluorine most reactive in periodic table
  - c. Group 1 will lose e- to become +1
  - d. Elements in same group have same chem properties.
3. CH 8: Ionic compounds and bonds
  - a. Why do atoms bond? How many electrons make them stable?
  - b. How to name ionic compounds, go from name to formula
  - c. What elements make up ionic compounds?
  - d. Properties of ionic compounds
  - e. What is a crystal lattice and the arrangement.
  - f. What is metallic bonding?
4. CH 9: Covalent compounds
  - a. What elements make up covalent compounds?
  - b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
  - c. What are pi bonds and sigma bonds?
  - d. Bond length Vs. Bond strength, and Bond dissociation energy.
  - e. What are molecules?
  - f. What are the common diatomic molecules?
  - g. Polar molecules, partial charges, what are they?
5. CH 10: Identify the types of chemical reactions
  - a. Be able to write the products for a double replacement and single replacement reactions.
  - b. Balancing equations
  - c. Writing equations from word names.
6. CH 11: The mole
  - a. Determine the molar mass of a compound
  - b. Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
  - c. Converting mole  $\rightarrow$  mole, mole  $\rightarrow$  grams, mole  $\rightarrow$  particles and visa versa
  - d. Converting grams  $\rightarrow$  mole, grams  $\rightarrow$  grams, grams  $\rightarrow$  particles visa vera
7. CH 12:
  - a. Stoichiometry conversions using balanced chemical equation
  - b. Mole  $\rightarrow$  mole, mole  $\rightarrow$  grams, grams  $\rightarrow$  grams

# Objective Final Review For Chemistry

**M Mosston**



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