

- (a) N_2 (b) O_2
 (c) Cl_2 (d) S

10. Which gas is used to manufacture king of chemicals sulphuric acid?

- (a) N_2 (b) O_2
 (c) Cl_2 (d) S

11. Equilibrium constant has no unit when number of moles of reactants and products are;

- (a) same (b) different
 (c) both a & b (d) none

12. For reactions having large K_c value, the reaction proceeds to;

- (a) completion
 (b) equilibrium state
 (c) back ward (d) None

13. The characteristics of reversible reactions are the following except;

- (a) products never recombine to form reactants
 (b) they never complete
 (c) they proceed in both ways
 (d) they have a double arrow between reactants and products

14. In the lime kiln, the reaction



goes to completion because;

- (a) of high temperature
 (b) CaO is more stable than $CaCO_3$
 (c) CO_2 escapes continuously
 (d) CaO is not dissociated

15. For the reaction,



the expression for the equilibrium constant is:

- (a) $\frac{[2A][B]}{[3C]}$ (b) $\frac{[A]^2[B]}{[C]^3}$

- (c) $\frac{[3C]}{[2A][B]}$ (d) $\frac{[C]^3}{[A]^2[B]}$

16. When a system is at equilibrium states?

- (a) the concentration of reactants and products becomes equal
 (b) the opposing reactions (forward and reverse) stop
 (c) the rate of the reverse reaction becomes very low
 (d) the rates of the forward and reverse reactions becomes equal.

17. Which one of the following statements is not correct about active mass?

- (a) rate of reaction is directly proportional to active mass.
 (b) active mass is taken in molar concentration
 (c) active mass is represented by square brackets
 (d) active mass means total mass of substances

18. When the magnitude of K_c is very large it indicates;

- (a) reaction mixture consists of almost all products
 (b) reaction mixture has almost all reactants
 (c) reaction has not gone to completion
 (d) reaction mixture has negligible products

19. When the magnitude of K_c is very small it indicates;

- (a) equilibrium will never establish

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Mark Steyvers



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