

Number

...or **NUMB**, for the correct order of operations, take care when using a calculator.

- Brackets
- Orders (or powers)
- Division and Multiplication
- Addition and Subtraction

Types of number

Integer: a 'whole' number
Factors: the divisors of an integer
• Factors of 12 are 1, 2, 3, 4, 6, 12
Multiples: a 'times table' for an integer (with infinite multiples)
• Multiples of 12 are 12, 24, 36, ...
Prime number: an integer which has exactly two factors (1 and the number itself). Note it is not a prime number.

Units

Highest Common Factor (HCF)
• Factors of 6 are 1, 2, 3, 6
Factors of 9 are 1, 3, 9
HCF of 6 and 9 is 3

Lowest Common Multiple (LCM)

• Multiples of 6 are 6, 12, 18, 24, ...
Multiples of 9 are 9, 18, 27, 36, ...
LCM of 6 and 9 is 18

Power notation

Write a number as a product of its prime factors, and follow for repeated factors.
• $120 = 2 \times 2 \times 2 \times 3 \times 5$

Indices and roots

Special indices for any number a
 $a^0 = 1$
 $a^{-1} = \frac{1}{a}$
 $a^{\frac{1}{2}} = \sqrt{a}$

Ordering with fractions

Adding or subtracting fractions, use a common denominator.
• $\frac{1}{2} + \frac{1}{3} = \frac{3}{6} + \frac{2}{6} = \frac{5}{6}$

Multiplying fractions

Multiplying fractions: multiply numerators and denominators.
• $\frac{1}{2} \times \frac{1}{3} = \frac{1 \times 1}{2 \times 3} = \frac{1}{6}$

Working fractions 'top' the second fraction

Working fractions 'top' the second fraction, then multiply.
• $\frac{1}{2} \div \frac{1}{3} = \frac{1}{2} \times \frac{3}{1} = \frac{3}{2}$

Proportion notation

Fraction in numerator = denominator
• $\frac{1}{2} \div \frac{1}{3} = \frac{1}{2} \times \frac{3}{1} = \frac{3}{2}$

Use the given values to change decimals or fractions, simplify where possible.

• $0.45 = \frac{45}{100} = \frac{9}{20}$

Leave the most frequently used ones

100	10	1	0.1	0.01	0.001
100	10	1	0.1	0.01	0.001

Units

Look for the biggest square number factor of the number.
• $100 = 2 \times 2 \times 5 \times 5 \times 5$

Standard form

Standard form numbers are of the form: $a \times 10^n$ where $1 \leq a < 10$ and n is an integer.

Scientific notation

1 atom = 0.000 000 1 kilograms
1 kilogram = 1 000 grams
1 kilometre = 1 000 metres
1 metre = 100 centimetres
= 1 000 millimetres
1 centimetre = 10 millimetres

1 day = 24 hours
1 hour = 60 minutes = 3 600 seconds
1 minute = 60 seconds

Area and perimeter

Calculate the number, then use it 'double right' to round up or down. Round (up) means use the decimal point.

• 100.1 rounds to 100
• 100.5 rounds to 101

100.51 rounds to 101
100.49 rounds to 100

100.50 rounds to 101
100.49 rounds to 100

100.51 rounds to 101
100.49 rounds to 100

Area and perimeter

Find the area of numbers that will round to a given value.
• $a = 5.55$ (2 decimal places)
 $5.55 \leq a < 5.56$

• $a = 5.55$ (2 significant figures)
 $5.5 \leq a < 5.6$

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Area and perimeter

Use the area of a square, and that the two opposite sides are equal.

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Standard graphs



Equation of a straight line

Equation of straight line $y = mx + c$ as in the graph, c is the y -intercept.
• Find the equation of the line that joins (0, 2) to (2, 1).
Find the gradient: $m = \frac{y_2 - y_1}{x_2 - x_1} = \frac{1 - 2}{2 - 0} = -\frac{1}{2}$
and the y -intercept, $c = 2$
Hence the equation is $y = -\frac{1}{2}x + 2$
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Equation of a straight line

Parallel lines: gradients are equal.
• $y = 3x + 2$ and $y = 3x + 4$ both have gradient 3 so are parallel.

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Right-angled triangles



Pythagorean Theorem

Pythagorean Theorem: In a right-angled triangle, the square of the hypotenuse is equal to the sum of the squares of the other two sides.
No angles: $a^2 + b^2 = c^2$

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Area and perimeter

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Area of rectangle = $\text{length} \times \text{width}$
Area of circle = πr^2

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VM Jensen



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Table of Contents Pixl May 2015 Ocr Style Higher Tier

1. Understanding the eBook Pixl May 2015 Ocr Style Higher Tier
 - The Rise of Digital Reading Pixl May 2015 Ocr Style Higher Tier
 - Advantages of eBooks Over Traditional Books
2. Identifying Pixl May 2015 Ocr Style Higher Tier
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pixl May 2015 Ocr Style Higher Tier
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pixl May 2015 Ocr Style Higher Tier
 - Personalized Recommendations
 - Pixl May 2015 Ocr Style Higher Tier User Reviews and Ratings
 - Pixl May 2015 Ocr Style Higher Tier and Bestseller Lists
5. Accessing Pixl May 2015 Ocr Style Higher Tier Free and Paid eBooks
 - Pixl May 2015 Ocr Style Higher Tier Public Domain eBooks
 - Pixl May 2015 Ocr Style Higher Tier eBook Subscription Services
 - Pixl May 2015 Ocr Style Higher Tier Budget-Friendly Options
6. Navigating Pixl May 2015 Ocr Style Higher Tier eBook Formats

- ePub, PDF, MOBI, and More
- Pixl May 2015 Ocr Style Higher Tier Compatibility with Devices
- Pixl May 2015 Ocr Style Higher Tier Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pixl May 2015 Ocr Style Higher Tier
 - Highlighting and Note-Taking Pixl May 2015 Ocr Style Higher Tier
 - Interactive Elements Pixl May 2015 Ocr Style Higher Tier
- 8. Staying Engaged with Pixl May 2015 Ocr Style Higher Tier
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pixl May 2015 Ocr Style Higher Tier
- 9. Balancing eBooks and Physical Books Pixl May 2015 Ocr Style Higher Tier
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pixl May 2015 Ocr Style Higher Tier
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pixl May 2015 Ocr Style Higher Tier
 - Setting Reading Goals Pixl May 2015 Ocr Style Higher Tier
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pixl May 2015 Ocr Style Higher Tier
 - Fact-Checking eBook Content of Pixl May 2015 Ocr Style Higher Tier
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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jun 23 2022 what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4^+ \text{aq} + 2\text{F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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