

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

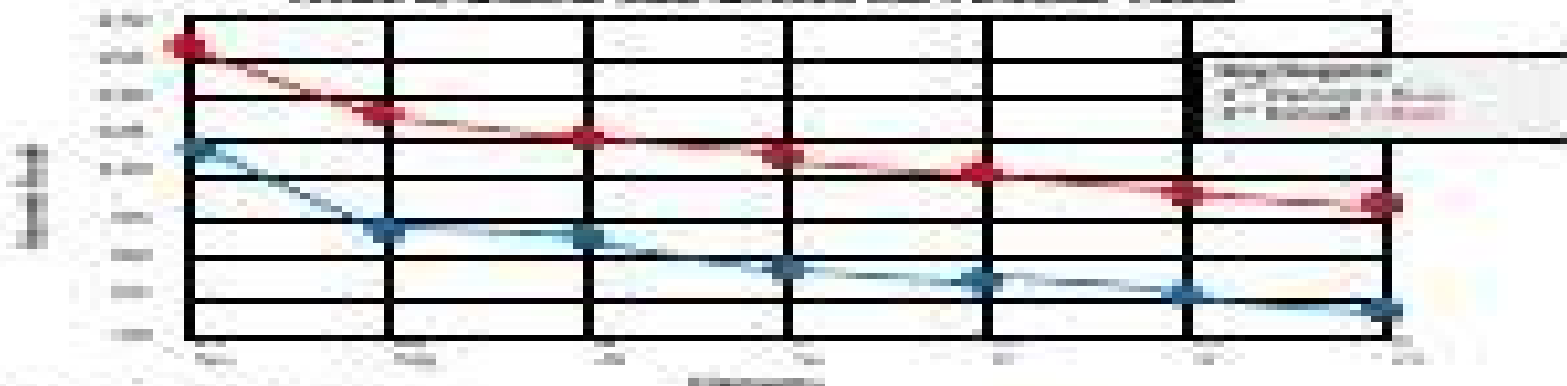
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property. Sketch it's data points graphically.
2. Is a slight downward trend in this graph Period? If not, fill in the other graphs.
3. Label the box as 'not periodic' if not periodic.
4. Answer all questions.

	1	2	3	4	5	6	7	8
1	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
2	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
3	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
4	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
5	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
6	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
7	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
8	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
9	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity
10	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity	Electronegativity

Now label the trends in the following way:

Trend in Atomic Size across the Periodic Table



Electronegativity trend questions

1. Following the trend in atomic size across the periodic table, decreases
2. This causes two more than just more electrons protons/electrons. They give all electrons, pushing them away, which is increasing the force.

Electronegativity trend questions

1. Electronegativity trend in the periodic table is increases (e.g. Na, Cl, Br, I, At). The trend in electronegativity increases as you go across the periodic table.
2. This causes two more than just more electrons protons/electrons (e.g. protons and electrons) pushing them away, increasing electronegativity.

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ML Yell



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