Period Date Tuesday, November 1 2016

Periodic Trends

ATOMIC RADIUS

- What trend in atomic radius do you see as you go down a group/family on the periodic table?
- What causes this trend?
- 3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
- 4. What causes this trend?
- Circle the atom in each pair that has the largest atomic radius.
 - a) Al B

- b) S O
- c) Br Cl

- d) Na Al e) O F f) Mo Ca
- Put the following elements in order from smallest to largest atomic radius and explain why: C. O. Sn. Sr.

ELECTRONEGATIVITY

- Define electronegativity
- 8. How does the ionic radius of a nonmetal compare with its atomic radius?
- 9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
- 10. What causes this trend?
- 11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
- 12. What causes this trend?
- Circle the atom in each pair that has the greater electronegativity.
- a) Ca Ga

- b) Li O c) Cl S d) Br As
- e) Ba Sr

GENERAL QUESTIONS

- 14. Which group tends to form +1 ions?
- 15. Which group tends to form +2 ions?
- 16. Which group tends to form -1 ions?
- 17. Which group tends not to form ions or react?

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Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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