

The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓
(Columns)

The elements in each group have the same number of valence electrons.

Diagram illustrating the periodic table layout and electron shells.

Groups:

- Group 1: (1 valence e⁻)
- Group 2: (2 valence e⁻)
- Group 3-10: (8 valence e⁻)
- Group 11: (1 valence e⁻)
- Group 12: (2 valence e⁻)
- Group 13-18: (8 valence e⁻)

Periods and Shells:

- 1st Period → (1 shell)
- 2nd Period → (2 shells)
- 3rd Period → (3 shells)
- 4th Period → (4 shells)

Periodic Table Structure:

Group 1 (1 valence e ⁻)	Group 2 (2 valence e ⁻)	Group 3-10 (8 valence e ⁻)	Group 11 (1 valence e ⁻)	Group 12 (2 valence e ⁻)	Group 13-18 (8 valence e ⁻)		
1 st Period 1 H Hydrogen					2 He Helium		
2 nd Period 3 Li Lithium	4 Be Beryllium	5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon
3 rd Period 11 Na Sodium	12 Mg Magnesium	13 Al Aluminum	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon
4 th Period 19 K Potassium	20 Ca Calcium	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton

Notes:

- Except Helium with 2 valence e⁻

Periods
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

Periodic Table Notes For Middle School

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