

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- Based on Moseley law: $\sqrt{f} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements; they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 58 - 71$

G_3 and P_6

actinoids = $2 = 90 - 103$

G_3 and P_6

Diagonal Relation;

Li → Be → B → similar
Na → Mg → Al → change ratio
size

Nomenclature;

| | | |
|----------|----------|---------------------|
| 0 → nil | 5 → pent | Suffix = turn |
| 1 → un | 6 → hex | 100 unnilpentium |
| 2 → bi | 7 → sept | |
| 3 → tri | 8 → oct | |
| 4 → quad | 9 → non | |

EFFECTIVE NUCLEAR CHARGE

$Z_{eff} = Z - \sigma$

Trends - Period → $Z_{eff} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{eff} \uparrow \downarrow$

Shielding → $s > p > d > f$

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff} .

IONISATION ENERGY

- Energy required to remove an electron from an isolated gaseous atom.



IE = IE₁ < IE₂ < IE₃ ...

Trends - $Z_{eff} \propto$ IE

Period - $Z_{eff} \uparrow =$ IE \uparrow

Group - $Z_{eff} \downarrow =$ IE \downarrow

$G_2 = B > Al > Ga > In$

$G_{14} = C > Si > Ge > Pb > Sn$

Period 2:-

$Li < Be < B < C < N < O < F < Ne$

ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covalent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



$$R_c = \frac{d}{2}$$

Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



$$R_m = \frac{d}{2}$$

Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



$$R_v = \frac{d}{2}$$

Trends; Atomic Radius $\propto \frac{1}{Z_{eff}}$

Period → $Z_{eff} \uparrow =$ Atomic Radius \downarrow

Group → $Z_{eff} \downarrow =$ Atomic Radius \uparrow

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

- Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{eg}H_2$ is -ve (exothermic)

Exception = $\begin{matrix} B \\ N \end{matrix}$ } $\Delta_{eg}H_2 = +ve$ due to stable configuration noble gas

Note - Successive electron gain enthalpy is always +ve.

Trends; $\Delta_{eg}H \propto Z_{eff}$

Period - $Z_{eff} \uparrow \Rightarrow \Delta_{eg}H \uparrow$

Group - $Z_{eff} \downarrow \Rightarrow \Delta_{eg}H \downarrow$

$G_2 = S > Se > Te > Po > O$
→ due to small size Reason!
→ e^-e^- repulsion.

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