

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-3)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant

Excess: All reactants except the limiting reagent will have some left over. Excess is the amount of reactant that is not completely consumed.

Remember: Limiting reagent will determine how much product is produced and by how much. The limiting reagent is also a reactant.



The ratio required is the composition for the limiting reagent.

2H ₂ (g)	20.00 grams
O ₂ (g)	32.00 grams
2H ₂ O(l)	36.00 grams
H ₂ O	36.00 grams

1. 20.00 g of each reactant will be used & react to form 36.00 g of product from 20.00 g of hydrogen. Excess oxygen will be left over. 20.00 g of hydrogen is the limiting reagent.



2. 20.00 g H₂(g) is the limiting reagent and 32.00 g O₂(g) is the excess reagent.

3. 36.00 g of H₂O(l) is the amount of product that can be produced from 20.00 g of H₂(g) and 32.00 g of O₂(g).



4. 36.00 g of H₂O(l) is the amount of product that can be produced from 20.00 g of H₂(g) and 32.00 g of O₂(g). 36.00 g of H₂O(l) is the amount of product that can be produced from 20.00 g of H₂(g) and 32.00 g of O₂(g).



Percent Yield And Limiting Reactants Quiz

Richard Bailey



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