

Percent Composition Practice 1.0 - Answer Key

Objective: - Calculate the percent composition of a compound.

Directions: Determine the percent composition of each of the compounds below. (Show all work!)

1. HCl

$$\text{H: } 1.00\text{g} \times 1 = 1.01\text{g} \quad \left(\frac{1.01\text{g}}{35.46\text{g}} \right) \times 100\% = \boxed{2.77\%}$$

$$\text{Cl: } 35.45\text{g} \times 1 = 35.45\text{g} \quad \left(\frac{35.45\text{g}}{35.46\text{g}} \right) \times 100\% = \boxed{97.23\%}$$

36.46g

2. KMnO_4

$$\text{K: } 39.10\text{g} \times 1 = 39.10\text{g} \quad \left(\frac{39.10\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{24.74\%}$$

$$\text{Mn: } 54.94\text{g} \times 1 = 54.94\text{g} \quad \left(\frac{54.94\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{34.76\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g} \quad \left(\frac{64.00\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{40.50\%}$$

158.04g

3. $\text{Mg(NO}_3)_2$

$$\text{Mg: } 24.31\text{g} \times 1 = 24.31\text{g} \quad \left(\frac{24.31\text{g}}{148.33\text{g}} \right) \times 100\% = \boxed{16.39\%}$$

$$\text{N: } 14.01\text{g} \times 2 = 28.02\text{g} \quad \left(\frac{28.02\text{g}}{148.33\text{g}} \right) \times 100\% = \boxed{18.89\%}$$

$$\text{O: } 16.00\text{g} \times 6 = 96.00\text{g} \quad \left(\frac{96.00\text{g}}{148.33\text{g}} \right) \times 100\% = \boxed{64.72\%}$$

148.33g

4. $(\text{NH}_4)_3\text{PO}_4$

$$\text{N: } 14.01\text{g} \times 3 = 42.03\text{g} \quad \left(\frac{42.03\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{28.19\%}$$

$$\text{H: } 1.01\text{g} \times 12 = 12.12\text{g} \quad \left(\frac{12.12\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{8.13\%}$$

$$\text{P: } 30.97\text{g} \times 1 = 30.97\text{g} \quad \left(\frac{30.97\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{20.77\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g} \quad \left(\frac{64.00\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{42.92\%}$$

149.12g

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