

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- math

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the collisions of the gas molecules impinging the surface.
- If the gas is compressed to smaller the volume it has more surface impacts the pressure increases with volume.
- Force exerts pressure on area of the walls of the container.
- Pressure of the gas will increase.

- Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. If the gas is released, it maintains a volume of 900 ml. above the pressure inside of 800 Pa. What is the volume of the balloon at this altitude?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 101.3 \text{ kPa} \quad P_2 = 80 \text{ kPa} \quad V_1 = 800 \text{ ml} \quad V_2 = ?$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{101.3 \times 800}{80} = 1013 \text{ ml}$$

- Ex. 2. A gas at pressure of 101.3 kPa and 400 ml. of volume. If the pressure is increased until the volume is 300 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 101.3 \text{ kPa} \quad V_1 = 400 \text{ ml} \quad V_2 = 300 \text{ ml} \quad P_2 = ?$$

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{101.3 \times 400}{300} = 135.1 \text{ kPa}$$

Note Taking Guide Episode 902 Chemistry

Liyong Dong



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Working Mother ,2002-10 The magazine that helps career moms balance their personal and professional lives

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