

## Trends on the Periodic Table

### Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

**Lithium**  
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



**Francium**  
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

**Fluorine**  
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



I am so perfect... I have a full valence shell. I don't need those other elements... I hope they just leave me alone!



**Neon**  
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

**Iodine**  
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

### Part 2: Use your textbook to complete the following:

#### Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

*Effective nuclear charge,  $Z_{eff}$ , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.*

*not a good enough answer*  
**4AP Answer**

# Periodic Trends Answers Key Pogil

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## Enjoying the Melody of Term: An Mental Symphony within **Periodic Trends Answers Key Pogil**

In a global eaten by screens and the ceaseless chatter of immediate communication, the melodic splendor and emotional symphony produced by the written word frequently fade into the backdrop, eclipsed by the persistent noise and interruptions that permeate our lives. However, situated within the pages of **Periodic Trends Answers Key Pogil** a marvelous literary treasure brimming with organic thoughts, lies an immersive symphony waiting to be embraced. Crafted by a wonderful musician of language, that fascinating masterpiece conducts visitors on a psychological trip, skillfully unraveling the concealed songs and profound influence resonating within each carefully constructed phrase. Within the depths of this emotional analysis, we will examine the book is main harmonies, analyze its enthralling publishing style, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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### **Table of Contents Periodic Trends Answers Key Pogil**

1. Understanding the eBook Periodic Trends Answers Key Pogil
  - The Rise of Digital Reading Periodic Trends Answers Key Pogil
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Answers Key Pogil
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Answers Key Pogil
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Answers Key Pogil
  - Personalized Recommendations

- Periodic Trends Answers Key Pogil User Reviews and Ratings
- Periodic Trends Answers Key Pogil and Bestseller Lists
- 5. Accessing Periodic Trends Answers Key Pogil Free and Paid eBooks
  - Periodic Trends Answers Key Pogil Public Domain eBooks
  - Periodic Trends Answers Key Pogil eBook Subscription Services
  - Periodic Trends Answers Key Pogil Budget-Friendly Options
- 6. Navigating Periodic Trends Answers Key Pogil eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Answers Key Pogil Compatibility with Devices
  - Periodic Trends Answers Key Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Answers Key Pogil
  - Highlighting and Note-Taking Periodic Trends Answers Key Pogil
  - Interactive Elements Periodic Trends Answers Key Pogil
- 8. Staying Engaged with Periodic Trends Answers Key Pogil
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Answers Key Pogil
- 9. Balancing eBooks and Physical Books Periodic Trends Answers Key Pogil
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Answers Key Pogil
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Answers Key Pogil
  - Setting Reading Goals Periodic Trends Answers Key Pogil
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Answers Key Pogil
  - Fact-Checking eBook Content of Periodic Trends Answers Key Pogil

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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