

Concepts in Chemistry

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the periodic law.

1. _____ for the modern periodic table, elements are ordered
 - a) according to increasing atomic mass.
 - b) according to Mendeleev's original design.
 - c) according to increasing atomic number.
 - d) based on how they were discovered.
2. _____ Mendeleev ordered the elements in his periodic table according to
 - a) properties appeared in regular intervals and observed chemical laws.
 - b) density.
 - c) atomic number.
 - d) mass.
3. _____ The modern periodic law states that
 - a) all new elements will be added upon as he found in the future group areas.
 - b) the physical and chemical properties of elements are functions of atomic number.
 - c) elements exhibit properties of both periods and groups.
 - d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot.
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - a) adding a new
 - b) period.
 - c) group.
 - d) series.
5. _____ The noble gases that property of the noble gases is that they are
 - a) unstable.
 - b) inert.
 - c) radioactive.
 - d) highly reactive.
6. _____ Likewise, the first element in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
 - a) 4.
 - b) 5.
 - c) 11.
 - d) 19.
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - a) 2. How many periods are in this series?
 - b) 3. How many elements are in this group?
 - c) 4. How many elements are in this group?
 - d) 5. How many elements are in this group?

Periodic Law Chapter 5 Test Answers

James E. Brady, Fred Senese



Periodic Law Chapter 5 Test Answers:

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