

Chapter 7.1

Solutions

- Homogeneous mixtures of two or more substances in which the components are present as atoms, molecules, or ions
- Particles in liquid solutions are:
 - Too small to reflect light, thus solutions are transparent
 - In constant motion
 - Not settled by the influence of gravity
- Some solutions are colored
- Solvent of a solution
 - The substance present in the largest amount of the solution
- Solute of a solution
 - Any substance present in an amount less than that of the solvent
 - A solution may contain more than one solute
- Characteristics of a liquid solution
 - It has uniform composition
 - It cannot be separated by filtration
 - It can be separated into pure components
- Physical states of solutions
 - The physical state of a solution of a solid, liquid, or gas, is usually the same as the physical state of the solvent

TABLE 7.1 Solutions in various states

Solution	Solution state	Solvent	Solute
Salt water	Liquid	Water (liquid)	Sodium chloride (solid)
Alcoholic beverage	Liquid	Water (liquid)	Alcohol (liquid)
Carbonated water	Liquid	Water (liquid)	Carbon dioxide (gas)
Gold alloy (jewelry)	Solid	Gold (solid)	Copper (solid)
Gold amalgam	Solid	Gold (solid)	Mercury (liquid)
Hydrogen in palladium	Solid	Palladium (solid)	Hydrogen (gas)
Air	Gaseous	Nitrogen (gas)	Oxygen (gas)
Mixed oxygen	Gaseous	Oxygen (gas)	Water (liquid)
Camphor in nitrogen	Gaseous	Nitrogen (gas)	Camphor (solid)

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- Solubility
 - The solubility of a solute is the maximum amount of the solute that can be dissolved in a specific amount of solvent under specific conditions of temperature and pressure
 - Solute solubility

Noun Chm 103

Mark Steyvers



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