

Chapter 12 – Stoichiometry

Note Taking Guide: Episode 801 Stoichiometry

-Study of the **QUANTITY** relationships in a **CHEMICAL REACTION**

-Based on **BALANCED** equations



1. Particles

2 atoms Mg	2 atoms O	2 atoms Mg
		2 atoms O

2 atoms Mg	1 molecule O ₂	2 molecule MgO
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10 atoms Mg	5 molecule O ₂	10 molecule MgO
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2 (6.02 x 10 ²³) atoms Mg	1 (6.02 x 10 ²³) molecule O ₂	2 (6.02 x 10 ²³) molecule MgO
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2. Moles

2 mol Mg	1 mol O ₂	2 mol MgO
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The **COEFFICIENTS** in a **BALANCED EQUATION** give the **MOLE RATIO** for the **SUBSTANCES** involved in the **REACTION**.

Lab: Episode 801 - Can You Find the Mole Ratio?

In this experiment, baking soda (NaHCO₃) reacts with hydrochloric acid (HCl) to produce sodium chloride, carbon dioxide, and water. You will determine the moles of reactant used and moles of product produced using both the lab data and stoichiometry and compare the two.

Data Table:

mass of test tube and baking soda	19.46 g
mass of empty test tube	19.21 g

Note Taking Guide Episode 801 Stoichiometry

Mark Steyvers



Note Taking Guide Episode 801 Stoichiometry:

Decoding **Note Taking Guide Episode 801 Stoichiometry**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Note Taking Guide Episode 801 Stoichiometry**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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