

Quantum Numbers

- Used to _____ an _____ in an _____

n

- _____
- Represents _____ energy level of _____
- _____ of _____ in an _____

Example: What is the maximum number of electrons that can be in the _____ main energy level?

l

- The _____
- Describes the _____ within an _____
- _____ of orbital _____ possible is _____

Orbital Shapes

designated _____

- level 1: _____
- level 2: _____
- level 3: _____
- level 4: _____

How many electrons can each sublevel hold?

s = 1 orbital $\times$ 2 e⁻/orbital = _____e⁻
 p = 3 orbitals $\times$ 2 e⁻/orbital = _____e⁻
 d = 5 orbitals $\times$ 2 e⁻/orbital = _____e⁻
 f = 7 orbitals $\times$ 2 e⁻/orbital = _____e⁻

Note Taking Guide Episode 304 Key

Paul H. D. Kaplan



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