

Part 2

GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

Combined Science Higher - Chemistry

Walkthrough Mark Paper

12
19

Look at the equation for the equilibrium reaction between sulfur dioxide, SO_2 , oxygen and sulfur trioxide, SO_3 , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass, A_r , of S is 32.1 and of O is 16.0.

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Revise A2 Chemistry for Salters (OCR) Daniels,2005-04-04 Helps students to pull together key ideas in the course and apply them to exam questions in a fresh context Organised by module to allow readers to quickly access specific information this work provides tips on common pitfalls and advice on approaching exam questions with practice style exam questions for each module along with answers

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OCR A Level Chemistry Student Book 2 Mike Smith,John Older,2015-09-11 Exam Board OCR Level A

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