

Nobel - OCR: Chemistry Unit 2 – Chains, Energy and Resources

(a) Interpret and use the terms:

(i) Empirical formula = simplest whole number ratio of atoms of each element present in a compound

(ii) Molecular formula = actual number of atoms of each element in a molecule,

(iii) General formula = simplest algebraic formula of a member of a homologous series

Alkane	C_nH_{2n+2}
Alkene	C_nH_{2n}
Alcohol	$C_nH_{2n+2}OH$
Cycloalkane	C_nH_{2n}
Cycloalkene	C_nH_{2n-2}
Haloalkane	$C_nH_{2n+2}X$

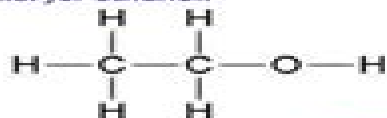
(iv) Structural formula = minimal detail that shows the arrangement of atoms in a molecule

E.g. Butane:

$CH_3CH_2CH_2CH_3$ or $CH_3(CH_2)_2CH_3$.

(v) Displayed formula = relative positioning of atoms and the bonds between them,

i.e. for ethanol:



(vi) Skeletal formula = simplified organic formula, shown by removing hydrogen atoms from alkyl chains, leaving just a carbon skeleton and associated functional group

ie for butan-1-ol:



(b) Interpret, and use, the terms:

(i) homologous series = a series of organic compounds having the same functional group but with each successive member differing by CH_2 .

(ii) functional group = a group of atoms responsible for the characteristic reactions of a compound;

(c) Use the general formula of a homologous series to predict the formula of any member of the series;

E.g – Alkane = C_nH_{2n+2}

If $C = 9$, $H = (9 \times 2) + 2$

Therefore = C_9H_{20} = Nonane

(d) State the names of the first ten members of the alkane homologous series;

No of Carbons	Alkyl Group
1	Methyl
2	Ethyl
3	Propyl
4	Butyl
5	Pentyl
6	Hexyl
7	Heptyl
8	Octyl
9	Nonyl
10	Decyl

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David Lees, John Payne



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