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# The Royal Marsden Manual of Clinical Nursing Procedures

Edited by Sara Lister, Justine Hofland and Hayley Grafton

**Tenth Edition**



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# Royal Marsden Manual

**Patrick Vollmar**



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid  $\text{H}_2\text{CO}_3$  and its conjugate base the bicarbonate ion  $\text{HCO}_3^-$

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arrhenius s definition of acids and bases the earliest definition of acids and bases is arrhenius s definition which states that

an acid is a substance that forms hydrogen ions  $H^+$  when dissolved in water and a base is a substance that forms hydroxide ions  $OH^-$  when dissolved in water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is

$$acid + base \rightarrow water + salt$$

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