

Calculus-Based Properties (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equilibrium is at 100°C and $P_g = 1$ atm.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{moles solute}}{\text{kg solvent}} = \frac{1.86^\circ\text{C} \times 5.56 \text{ mol/kg}}{1} = 10.3^\circ\text{C}$$

$\Delta T_f = 1.86^\circ\text{C} \times 5.56 \text{ mol/kg} = 10.3^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 10.3^\circ\text{C} = -10.3^\circ\text{C}$
 $\Delta T_b = 0.51^\circ\text{C} \times 5.56 \text{ mol/kg} = 2.84^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 2.84^\circ\text{C} = 102.84^\circ\text{C}$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equilibrium is at 100°C and $P_g = 1$ atm. $K_f = 1.86^\circ\text{C}$, $K_b = 0.51^\circ\text{C}$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60 \text{ g}}}{0.100 \text{ kg}} = 16.7 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{moles solute}}{\text{kg solvent}} = \frac{1.86^\circ\text{C} \times 16.7 \text{ mol/kg}}{1} = 31.1^\circ\text{C}$$

$\Delta T_f = 1.86^\circ\text{C} \times 16.7 \text{ mol/kg} = 31.1^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 31.1^\circ\text{C} = -31.1^\circ\text{C}$
 $\Delta T_b = 0.51^\circ\text{C} \times 16.7 \text{ mol/kg} = 8.52^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 8.52^\circ\text{C} = 108.52^\circ\text{C}$

3. A 100 g mass percent (100 g $C_6H_{12}O_6$, 100 g H_2O) solution is at 100°C and $P_g = 1$ atm. Assume that the concentration of the solution is 100 g $C_6H_{12}O_6$ / 100 g H_2O . Assume that the concentration of the solution is 100 g $C_6H_{12}O_6$ / 100 g H_2O .

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = \frac{K_f \times \text{moles solute}}{\text{kg solvent}} = \frac{1.86^\circ\text{C} \times 5.56 \text{ mol/kg}}{1} = 10.3^\circ\text{C}$$

$$T_f = 0^\circ\text{C} - 10.3^\circ\text{C} = -10.3^\circ\text{C}$$

$$T_b = 100^\circ\text{C}$$

$$\Delta T_b = \frac{K_b \times \text{moles solute}}{\text{kg solvent}} = \frac{0.51^\circ\text{C} \times 5.56 \text{ mol/kg}}{1} = 2.84^\circ\text{C}$$

$$T_b = 100^\circ\text{C} + 2.84^\circ\text{C} = 102.84^\circ\text{C}$$

$$\Delta T_b = 0.51^\circ\text{C} \times 5.56 \text{ mol/kg} = 2.84^\circ\text{C}$$

$$\Delta T_b = \frac{K_b \times \text{moles solute}}{\text{kg solvent}} = \frac{0.51^\circ\text{C} \times 5.56 \text{ mol/kg}}{1} = 2.84^\circ\text{C}$$

4. A solution composed from 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water is composed. The freezing-point of this solution is -10.3°C . The normal boiling-point of water is 100°C and the freezing-point of pure water is 0°C . Determine the molal concentration of the compound. Assume $K_f = 1.86^\circ\text{C}$, $K_b = 0.51^\circ\text{C}$.

$$\Delta T_f = \frac{K_f \times \text{moles solute}}{\text{kg solvent}} = 10.3^\circ\text{C} = \frac{1.86^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}} \Rightarrow \text{moles solute} = 5.56 \text{ mol}$$

$$\Delta T_f = 10.3^\circ\text{C} = \frac{K_f \times \text{moles solute}}{\text{kg solvent}}$$

$$10.3^\circ\text{C} = \frac{1.86^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}}$$

$$\text{moles solute} = 5.56 \text{ mol}$$

$$\Delta T_b = \frac{K_b \times \text{moles solute}}{\text{kg solvent}} = 2.84^\circ\text{C} = \frac{0.51^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}}$$

$$2.84^\circ\text{C} = \frac{0.51^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}}$$

$$\text{moles solute} = 5.56 \text{ mol}$$

$$\Delta T_b = 2.84^\circ\text{C} = \frac{0.51^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}}$$

$$\Delta T_b = \frac{K_b \times \text{moles solute}}{\text{kg solvent}} = 2.84^\circ\text{C} = \frac{0.51^\circ\text{C} \times \text{moles solute}}{0.100 \text{ kg}}$$

$$\text{moles solute} = 5.56 \text{ mol}$$

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Padhraic Smyth



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