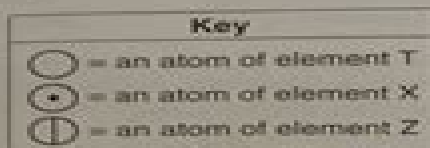


39 Given four particle models:



I



II



III



IV

Which two models can be classified as elements?

- (1) I and II
(2) I and IV
(3) II and III
(4) II and IV

40 After being thoroughly stirred at $10.^{\circ}\text{C}$, which mixture is heterogeneous?

- (1) 25.0 g of KCl and 100. g of H_2O unsat.
(2) 25.0 g of KNO_3 and 100. g of H_2O supersat.
(3) 25.0 g of NaCl and 100. g of H_2O unsat.
(4) 25.0 g of NaNO_3 and 100. g of H_2O unsat.

41 Which two compounds are electrolytes?

- (1) KOH and CH_3COOH
(2) KOH and C_3H_{12}
(3) CH_3OH and CH_3COOH
(4) CH_3OH and C_3H_{12}

42 Which statement explains why a CO_2 molecule is nonpolar?

- (1) Carbon and oxygen are both nonmetals.
(2) Carbon and oxygen have different electronegativities.
(3) The molecule has a symmetrical distribution of charge.
(4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

- (1) $15.^{\circ}\text{C}$ to 298 K $\uparrow$
(2) $37.^{\circ}\text{C}$ to 273 K $\downarrow$
(3) 305 K to $0.^{\circ}\text{C}$ $\downarrow$
(4) 355 K to $25.^{\circ}\text{C}$ $\downarrow$

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? Solid

- (1) N_2 (g)
(2) H_2 (g)
(3) Mg (s)
(4) Kr (g)

45 Which type of equilibrium exists in a sealed flask containing $\text{Br}_2(\ell)$ and $\text{Br}_2(\text{g})$ at 298 K and 1.0 atm ?

- (1) static phase equilibrium
(2) static solution equilibrium
(3) dynamic phase equilibrium
(4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

- (1) $\text{KH}(\text{s})$, $\text{Cl}^+(\text{aq})$, and $\text{OH}^-(\text{aq})$
(2) $\text{K}(\text{s})$, $\text{Cl}_2(\text{g})$, and $\text{H}_2\text{O}(\ell)$
(3) $\text{KCl}(\text{aq})$ and $\text{H}_2\text{O}(\ell)$
(4) $\text{KOH}(\text{aq})$ and $\text{Cl}_2(\text{g})$

47 In a titration, 20.0 milliliters of a 0.150 M $\text{NaOH}(\text{aq})$ solution exactly neutralizes 24.0 milliliters of an $\text{HCl}(\text{aq})$ solution. What is the concentration of the $\text{HCl}(\text{aq})$ solution?

- (1) 0.125 M
(2) 0.180 M
(3) 0.250 M
(4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

- (1) $\frac{1}{2}$
(2) $\frac{1}{3}$
(3) $\frac{1}{4}$
(4) $\frac{1}{8}$

Handwritten calculation:

$$3 = \frac{87.3}{29.14}$$

$$\frac{1}{2} = \frac{1}{8}$$

Regents Chemistry Topic 4 Review Answers

Albert S. Tarendash



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1. Understanding the eBook Regents Chemistry Topic 4 Review Answers
 - The Rise of Digital Reading Regents Chemistry Topic 4 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Regents Chemistry Topic 4 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Regents Chemistry Topic 4 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Regents Chemistry Topic 4 Review Answers
 - Personalized Recommendations
 - Regents Chemistry Topic 4 Review Answers User Reviews and Ratings
 - Regents Chemistry Topic 4 Review Answers and Bestseller Lists
5. Accessing Regents Chemistry Topic 4 Review Answers Free and Paid eBooks
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 - Regents Chemistry Topic 4 Review Answers eBook Subscription Services
 - Regents Chemistry Topic 4 Review Answers Budget-Friendly Options
6. Navigating Regents Chemistry Topic 4 Review Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Regents Chemistry Topic 4 Review Answers Compatibility with Devices
 - Regents Chemistry Topic 4 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Regents Chemistry Topic 4 Review Answers
 - Highlighting and Note-Taking Regents Chemistry Topic 4 Review Answers
 - Interactive Elements Regents Chemistry Topic 4 Review Answers
8. Staying Engaged with Regents Chemistry Topic 4 Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Regents Chemistry Topic 4 Review Answers
9. Balancing eBooks and Physical Books Regents Chemistry Topic 4 Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Regents Chemistry Topic 4 Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Regents Chemistry Topic 4 Review Answers
- Setting Reading Goals Regents Chemistry Topic 4 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Regents Chemistry Topic 4 Review Answers
- Fact-Checking eBook Content of Regents Chemistry Topic 4 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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