Question	Answer	Marks	Guidance
number			A STATE OF THE STA
6 (c)	$M(NaNO_3) = 85.0 \text{ g mol}^{-1}$	B1	
	n(NaNO <sub>3</sub> ) = 0.04 mol	81	
	$n(O_2) = 0.02 \text{ mol}$	B1	
	Volume of O <sub>2</sub> = 0.0200 × 24 000 = 480 cm <sup>3</sup>	B1	
7 (a)	0.0250 × 23.0 = 0.575 g	81	
7 (b)	2Na(s) + 2H <sub>2</sub> O(l) → 2NaOH(aq) + H <sub>2</sub> (g)	81	
7 (c)	n(H <sub>2</sub> ) = 0.0125 mol	B1	
	Volume H <sub>2</sub> = 0.0125 × 24 000 = 300 cm <sup>3</sup>	B1	
7 (d) (i)	$n(NaOH) = n(Na) = 0.0250 = c \times 50/1000$ $c = 0.500 \text{ mol dm}^{-3}$	B1	
7 (d) (ii)	$c = 0.500 \times 40.0 = 20.0 \mathrm{g}\mathrm{dm}^{-3}$	B1	
8 (a)	58.5 g	B1	
8 (b)	$n(Cl_2) = (2.5 \times 10^5)/24 = 1.04 \times 10^5 \text{ mol}$	B1	
	$n(NaOH) = 2 \times 1.04 \times 10^{8} = 2.08 \times 10^{8} \text{ mol}$	B1	
	$2.08 \times 10^8 = 4.00 \times V \text{ (in dm}^3\text{)}$ $\therefore V = 5.2 \times 10^7 \text{ dm}^3$	B1	
9 (a) (i)	C: H = 54.55/12.0: 9.09/1.0: 36.36/16.0 = 4.55 : 9.09: 2.27	B1	
	Empirical formula ≈ C <sub>2</sub> H <sub>4</sub> O	B1	
9 (a) (ii)	pV = nRT	B1	
	$n = \frac{(103 \times 10^3) \times (72.0 \times 10^{-6})}{8.314 \times 373} =$	B1	
	0.002 39 mol	B1	
	$M = \frac{0.2103}{0.00239} = 88.0$	B1	
9 (a) (iii)	Molecular formula = $C_2H_4O \times 88/44 = C_4H_8O_2$	B1	
10 (a) (i)	Al <sup>31</sup>	81	
	SO <sub>4</sub> 2-	B1	

# **Ocr Chemistry As Answers**

**Nora Henry** 

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