

## Lab 6: CHEMICAL EQUILIBRIUM LAB

**Introduction:** In this experiment you will start with a reaction that is at equilibrium. You will then change the concentrations of various ions that are present and note the effect on the state of equilibrium. The equilibrium reaction chosen has a colored product, so changes in its concentration can be readily detected by a corresponding change in color intensity.

### **Procedure:**

1. Pour about 25 mL of 0.0020 M potassium thiocyanate solution, KSCN, and 25 mL of distilled water into a beaker.
  - a. What is the color of this solution? \_\_\_\_\_
  - b. What are the 2 ions in this solution? \_\_\_\_\_
2. Look at the bottle of 0.20M iron (III)nitrate solution,  $\text{Fe}(\text{NO}_3)_3$  at your lab station.
  - a. What is the color of this solution? \_\_\_\_\_
  - b. What are the 2 ions in this solution? \_\_\_\_\_
3. Add 5 drops of the  $\text{Fe}(\text{NO}_3)_3$  solution to the beaker containing the potassium thiocyanate solution, KSCN. Gently swirl the mixture. What is the color of this solution? \_\_\_\_\_
4. There are now 4 ions in this solution. What are the two combinations of ions that could account for the color in step three?  
\_\_\_\_\_
5. At your lab station there is a solution of potassium nitrate  $\text{KNO}_3$ .  
What is the color of this solution? \_\_\_\_\_  
What are the two ions in this solution? \_\_\_\_\_
6. Based on your observations in step 5 and your answer to question 4, what are the two ions that must be responsible for the color you observed in step 5?  
\_\_\_\_\_

# Observing Chemical Equilibrium Lab 46 Answers

**Ernest Edgar Thum**



## Observing Chemical Equilibrium Lab 46 Answers:

**Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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**International Aerospace Abstracts** ,1999 **Chemical News and Journal of Industrial Science** ,1895

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**Colleges** William Crowell Bray, Wendell Mitchell Latimer, 1921      **The Chemical News and Journal of Physical Science**, 1895      **Plant Nutrition – from Genetic Engineering to Field Practice** J. Barrow, 2012-12-06 Plant Nutrition From Genetic Engineering to Field Practice the 12th International Colloquium on Plant Nutrition is the latest in a series which began in 1954 Early meetings were mainly concerned with the practical problems of soil fertility with soil assessment fertilizer requirements and methods of analysis As the colloquia have progressed the emphasis has slowly changed The practical problems are still important but there is increasing emphasis on plant physiology plant biochemistry membrane biochemistry and even on the chemistry of genes which control the proteins which transfer nutrient ions to the inside of cells The meetings therefore provide a valuable opportunity for each half of the science of plant nutrition to interact with and learn from the other half This volume begins with five papers which review current knowledge in important fields the rhizosphere molecular biology electron microscopy location and function of elements in vivo and modelling nutrient responses in the field These themes are continued in groups of shorter papers which follow In addition there are sections on nutrient dynamics and partitioning diagnostic techniques plant survival strategies mycorrhizas and on nutrients such as P N S K Ca Mg and micronutrients A large section is devoted specifically to boron reflecting the considerable current interest in this element In total there are 177 refereed papers providing both a broad overview and a detailed picture of the latest developments in pure and applied plant nutrition      **Technical Information Pilot**, 1949      *The Iron(III) Thiocyanate Reaction* Kevin C. de Berg, 2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book's main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation      *Techniques of Chemistry V 08 Prt 2 Dack Solutions* Weissberger, 1976

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