#### CHEMISTRY Objective Final Exam Review 10-11

There are 50 questions on the test, The chapters are listed below. Each of the letters represents a concept/idea that you should know or look up.

- 1. CH 6: Periodic trends, where do they increase & decrease
  - a. Atomic radius & Ionic radius
  - b. Electronegativity
  - c. Electron affinity
  - d. Ionization energy
  - e. Reactivity of metals and of non-metals
- 2. CH 7: Look up properties of elements per group and the names of each group.
  - Group 1 very reactive in water; group 7 most reactive non-metals etc.
  - b. Fluorine most reactive in periodic table
  - c. Group 1 will lose e- to become +1
  - d. Elements in same group have same chem properties.
- 3. CH 8: Ionic compounds and bonds
  - a. Why do atoms bond? How many electrons make them stable?
  - b. How to name ionic compounds, go from name to formula
  - c. What elements make up ionic compounds?
  - d. Properties of ionic compounds
  - e. What is a crystal lattice and the arrangement.
  - f. What is metallic bonding?
- 4. CH 9: Covalent compounds
  - a. What elements make up covalent compounds?
  - b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
  - c. What are pi bonds and sigma bonds?
  - d. Bond length Vs. Bond strength, and Bond dissociation energy.
  - e. What are molecules?
  - f. What are the common diatomic molecules?
  - g. Polar molecules, partial charges, what are they?
- 5. CH 10: Identify the types of chemical reactions
  - a. Be able to write the products for a double replacement and single replacement reactions.
  - b. Balancing equations
  - c. Writing equations from word names.
- 6. CH 11: The mole
  - a. Determine the molar mass of a compound
  - Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
  - c. Converting mole→ mole, mole → grams, mole → particles and visa versa
- d. Converting grams → mole, grams → grams, grams → particles visa vera
- 7. CH 12:
  - a. Stoichiometry conversions using balanced chemical equation
  - b. Mole → mole, mole → grams, grams → grams

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**Michal Rosen-Zvi** 

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