

10. Circle the element in each of the following pairs with the greater (1st) ionization energy:

a) lithium or boron

c) cesium or lead

b) magnesium or strontium

d) sodium or argon

11. Arrange the following elements in order of increasing ionization energy:

a) Sr, Be, Mg

b) Bi, Cs, Ba

c) Na, Al, S

12. Why is there a very large increase between the 1st and 2nd ionization energies for the elements in group 1?

Name: _____ Date: _____ Class period: _____



Unit: Periodic Trends "Ionic Size and E.N. Trend" - WS #4

Directions: Please answer each fill in the blank with the best answer. ☺

1. A cation is an ion with an overall _____ charge.
2. Cations are made when an atom _____.
3. The size of a positive ion is always _____ than the size of the neutral atom of the same element.
4. As the positive charge on an ion increases, the size of the atom will _____.
5. An anion is an ion with an overall _____ charge.
6. Anions are made when at atom _____.
7. The size of a negative ion is always _____ than the size of the neutral atom of the same element.
8. As the negative charge on an ion increases, the size of the atom will _____.
9. Which element/ion has the **biggest** atomic/ionic size in each of the following pairs? Please circle your answer.
 - a) Na or Na⁺¹
 - b) S or S⁻²
 - c) O⁻¹ or O⁻²
 - d) Al or Al⁺³
 - e) Cu⁺¹ or Cu⁺²
10. _____ is the tendency for the atoms of the element to attract electrons when they are chemically combined with another element.

Periodic Trends Unit Exam And Answers

RJ Shavelson



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