# Graphing Periodic Trends Lab

Name\_\_\_\_\_Hr\_\_\_

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

#### REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

ionization energy - The energy required to remove an electron from an atom.

#### DATA

Properties of the First 18 Elements

Properties of the First 18			erements	
Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy	
н	1	0.79	314	
He	2	0.49	567	
Li	3	2.05	124	
Ве	4	1.4	215	
8	5	1.17	191	
C	6	0.91	260	
M	7	0.75	335	
0	8	0.65	314	
F	9	0.57	402	
No	10	7	2	
Ma	1.1	2.23	119	
Mg	12	1.72	176	
A.I	13	1.82	138	
Si	1.4	1.46	188	
P	15	1.23	242	
S	16	1.09	239	
CI	17	0.97	299	
Ar	18	?	2	

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with guestions marks.

- Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)
- 2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)

# **Periodic Trends Lab Answers**

**David Baud** 

#### **Periodic Trends Lab Answers:**

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