

SAMANCOR BURSARY SCHEME 2015 APPLICATION FORM

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reverse reaction we learned about rates in unit five as how quickly a reaction proceeds

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web equilibrium questions a covalent bond has a standard gibbs free energy of 400 kJ mol while the hydrogen bond and van der Waals forces have substantially lower values respectively 25 and 5 kJ mol given such ΔG values which of the following is equivalent to the ratio of the K_{eq} value of a covalent bond to that of a hydrogen bond

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web in a chemical equilibrium the forward and reverse reactions occur at equal rates and the concentrations of products and reactants remain constant if we run a reaction in a closed system so that the products cannot escape we often find the reaction does not give a 100% yield of products