CLO- Classification of Elements & Periodicity in Properties MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

1.	Identify the least stable ion amongst the following :					
	(a) Li	(b) Be ⁻			(d) C	
2.	The correct order of radii is :					
	(a) N <be< b<="" td=""><td>(b) F < O2 < N3</td><td>(c) N< Li<k< td=""><td>(d) Fe</td><td>3*<fe3*<fe**< td=""></fe3*<fe**<></td></k<></td></be<>	(b) F < O2 < N3	(c) N< Li <k< td=""><td>(d) Fe</td><td>3*<fe3*<fe**< td=""></fe3*<fe**<></td></k<>	(d) Fe	3* <fe3*<fe**< td=""></fe3*<fe**<>	
3.	Which one of the following has largest size :					
	(a) Na	(b) Na*	(c) N	a	(d) Li*	
4.	The number of elements in the 4 th period of periodic table is :					
	(a) 8	(b) 10		В	(d) 32	
5.	Which of the following has least electron affinity (electron gain enthalpy):					
	(a) O	(b) N		r	(d) F	
6.	An element of atomic number 29 belongs to which of the following block of the periodic table :					
	(a) 5-block	(b) P-block	(c) d	-block	(d) f-block	
7.	Which one of the following arrangements not truly represents the property indicated against it:					
	(a) Br ₂ <cl<sub>2<f<sub>2 - electronegativity</f<sub></cl<sub>		(b) B	(b) Br ₂ <f<sub>2<cl<sub>2 - Electron Affinity</cl<sub></f<sub>		
	(c) Br ₂ <cl<sub>2 < F₂ - Bond energy</cl<sub>		(d) B	(d) Br ₂ <cl<sub>2<f<sub>2 - Oxidising Power</f<sub></cl<sub>		
8.	Alkali metals in each period have :					
	(a) Smallest size	(b) Lov	vest IE	(c) Highest IE	(d) Highest EN	
9.	A sudden large jump between the value of second and third ionization energies of elements would be					
	associated with which of the following electronic configuration :					
	(a) 2P*351	(b) 35 ²	3P1	(c) 35°3P°	(d) 2P°3S ²	
10.	Which of the following pairs are chemically dissimilar :					
	(a) Na and K	(b) Ba and Sr	(c) Z)	and Hf	(d) Ca and Zn	
111.	Collective name given to the elements with outer shell configuration ns np* is :					
	(a) Chalcogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases					
12.	Modern extended periodic table was given by :					
	(a) Bohr	(b) Lot	her mayer	(c) Moseley	(d) None alone	
13.	The most electronegative elements is :					
	(a) Hydrogen	(b) Fluorine	(c) C	hlorine	(d) Oxygen	

(a) 5-block (b) P-block

14. The element having electronic configuration : [Kr]4d104f14 552 5P4 5d1 652 belongs to

(c) d-block (d) f-block

 An elements R forms the highest oxide R₂O₁ R belong to : This study source was downloaded by 100000844331129 from CourseHero.com on 01-21-2024 10:32:22 GMT-06:00

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- 1. Understanding the eBook Periodicity Multiple Choice Questions
 - The Rise of Digital Reading Periodicity Multiple Choice Questions
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Periodicity Multiple Choice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodicity Multiple Choice Questions
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodicity Multiple Choice Questions
 - Personalized Recommendations
 - Periodicity Multiple Choice Questions User Reviews and Ratings
 - Periodicity Multiple Choice Questions and Bestseller Lists
- 5. Accessing Periodicity Multiple Choice Questions Free and Paid eBooks
 - Periodicity Multiple Choice Questions Public Domain eBooks
 - Periodicity Multiple Choice Questions eBook Subscription Services
 - Periodicity Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Periodicity Multiple Choice Questions eBook Formats

- o ePub, PDF, MOBI, and More
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- Periodicity Multiple Choice Questions Enhanced eBook Features
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 - o Adjustable Fonts and Text Sizes of Periodicity Multiple Choice Questions
 - Highlighting and Note-Taking Periodicity Multiple Choice Questions
 - Interactive Elements Periodicity Multiple Choice Questions
- 8. Staying Engaged with Periodicity Multiple Choice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Periodicity Multiple Choice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Multiple Choice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Multiple Choice Questions
 - Setting Reading Goals Periodicity Multiple Choice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Multiple Choice Questions
 - Fact-Checking eBook Content of Periodicity Multiple Choice Questions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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