

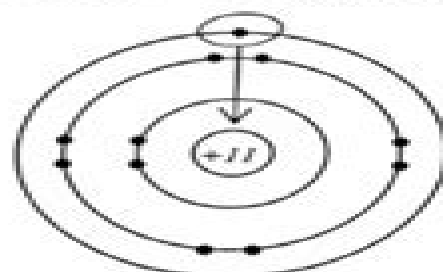
Advanced Periodic Trends

How do the particles in an atom interact to dictate periodic properties of elements?

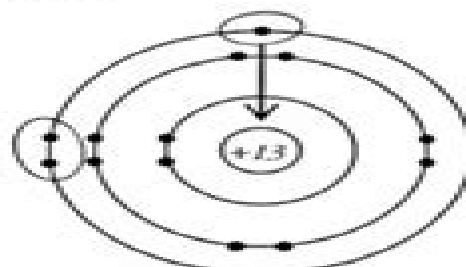
Why?

Previously you have learned about Coulombic attraction and how it governs the trends in properties among the elements of the Periodic Table. For example, as you move across a period, atoms have more protons in the nucleus, which pulls the electrons in tighter making smaller, more electronegative atoms. As you proceed down a column of the periodic table, the valence electrons get further away, and the atoms get larger and less electronegative. Inquisitive students will always ask, "If more protons in the nucleus pull electrons in tighter, then why don't atoms get smaller when going down a column?" That is a good question, and one that you will investigate in this activity.

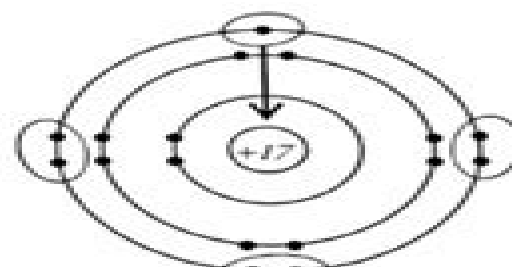
Model 1 – Period 3 Elements



Sodium



Aluminum



Chlorine

Atomic Number	11	13	17
Number of Electrons	11	13	17
Number of Valence Electrons	1	3	7
Core Charge	+1	+3	+7
Atomic Radius	186 pm	143 pm	99 pm
1st Ionization Energy	496 kJ/mol	578 kJ/mol	1251 kJ/mol
Electronegativity	0.9	1.5	3.0

1. Describe the relationship of the three elements in Model 1 with regard to their relative positions on the periodic table.

The three elements are in the same row of the periodic table.

2. Refer to a periodic table to complete the first three rows of the table in Model 1.

See Model 1.

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