

Gas
Liquid
Solid
Unknown

Periodic Table

Cheat Sheet

Reactions

Single Displacement $\rightarrow A + BC = B + AC$
 Double Displacement $\rightarrow AB + CD = BC + AD$
 Synthesis $\rightarrow A + B = AB$
 Decomposition $\rightarrow AB = A + B$
 Combustion $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$
 Neutralization $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$

Polyatomic ions (ion composed of multiple atoms)

$NO_3^- \rightarrow$ Nitrate (-1)

$PO_4^{3-} \rightarrow$ Phosphate (-3)

$OH^- \rightarrow$ Hydroxide (-1)

$HCO_3^- \rightarrow$ Bicarbonate (-1)

$NO_2^- \rightarrow$ Nitrite (-1)

$SO_4^{2-} \rightarrow$ Sulfate (-2)

$CO_3^{2-} \rightarrow$ Carbonate (-2)

$NH_4^+ \rightarrow$ Ammonium (+1)

Solid		Periodic Table										Liquid						Gas					
Unknown		Cheat Sheet										Decomposition $\rightarrow A + B = A + B$						Combustion $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$					
												Neutralization $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$											
Polyatomic ions (ion composed of multiple atoms)																							
$NO_3 \rightarrow \text{Nitrate } (-1)$																							
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$OH \rightarrow \text{Hydroxide } (-1)$																							
$HCO_3 \rightarrow \text{Bicarbonate } (-1)$																							
$NO_2 \rightarrow \text{Nitrite } (-1)$																							
$SO_4 \rightarrow \text{Sulfate } (-2)$																							
$CO_3 \rightarrow \text{Carbonate } (-2)$																							
$NH_4 \rightarrow \text{Ammonium } (+1)$																							

Alkali Metals:
 One valence electron
 Very reactive because they are ready to lose an electron to form cations
 Usually in ionic bonds due to reactivity
 Bonds with halogens, and other non-metals
 Have a vigorous reaction when bonding with halogens, oxygen and the compound of water

Alkaline Earth Metals:
 Two valence electrons
 Inactive but not as reactive as halogens and alkali metals because it needs to lose 2 electrons to have a full valence shell
 Bonds with halogens, and other non-metals

Transition Metals/ Metals: Transition metals are metals that vary in valence electrons and can have multiple ion charges. They have similar qualities as alkali metals and alkaline earth metals; they can form metallic and ionic bonds. Metals with multiple ion charges are called multivalent metals. For example, copper has an ion charge of +1 and +2 so it is a multivalent metal. Therefore, the ion charge must be specified when written in word form for ionic bonds with multivalent metals. Metals can also bond with polyatomic ions.

La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium
Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium

Metalloids: Contain metal and non-metal properties. They react with metals and non-metals to form covalent and ionic bonds but are less reactive than non-metals, depending on the electronegativity of the atoms.

Non-metals: They can not conduct electricity because they can not give off electrons, so current will not be able to go through. Non-metals bond with metals in an ionic bond and other non-metals in a molecular bond. They can also bond with the polyatomic ion ammonium due to its +1 charge.

Noble Gases: Non-metals in a previous form that have a full valence shell. This makes it not reactive, so it can not bond with other non-metals and metals. This is why noble gases are never in compounds and chemical reactions.

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D Keegan



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