

The Periodic Law

The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869. The amazing accuracy of his predictions has been very important to chemists in this century. However, the basis of his arrangement was the atomic masses of the elements. This approach proved incorrect as it would have placed some elements in a family with dissimilar properties. Henry Moseley rearranged the table on the bases of the atomic numbers of the elements. In accordance with Moseley's revision, the periodic law states: the properties of the elements are periodic functions of their atomic numbers.

Each of the 109 known elements has its own set of characteristic properties. These range from solid to gas, lustrous to dull, low to high melting point, various colors, and so on. The elements are arranged within the periodic table into groups or families (vertical columns) and periods or rows (horizontal rows). The arrangement reflects the periodic or repeating nature of the properties of the elements.

In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the elements from a scrambled periodic table. You will also predict values for any information missing from the table.

Objectives

In this exercise, you will

- Arrange the elements in Groups 1, 2, 13 – 18 (IA – VIIIA) according to a list of clues and your knowledge of periodic properties.
- Predict the missing properties of each element based on location in the table, and
- Explain the trends of properties in families and periods.

Procedure

1. Each block on the blank periodic table represents a different element from Groups 1 – 18 (IA – VIIIA).
2. Cut out blocks A – Z. Use the following clues to arrange the elements in their proper order on the blank periodic table. After you have placed these 26 elements in their correct position, glue them in place.
3. After you have placed blocks A – Z, cut out the remaining 16 blocks (*). Use the information provided in each block and your knowledge of periodic properties to arrange these elements in their proper position on the blank periodic table. Glue these blocks into place.
4. Some of the blocks have missing information. Predict the values for the missing items from the location of the element on the periodic table.

Periodic Law Lab Answers

RD Boyd



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