

POGIL: Periodic Trends in Atomic Properties

Information

Properties such as the size of an atom (atomic radius), the energy required to remove an electron from an atom (ionization energy), and the energy associated with the addition of an electron to a gaseous atom (electron affinity) can be understood in terms of the electron configuration of the atom and a competition between *electron-nucleus attraction* and *electron-electron repulsion*.

Preliminary Activity - Complete the orbital diagrams for the ground state configuration.

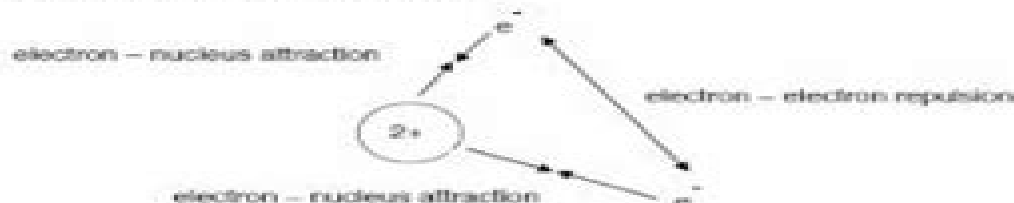
	1s	2s	2p		1s	2s	2p
H				C			
He				N			
Li				O			
Be				F			
B				Ne			

Model 1: Electron Ionization & Capture and Electrostatic Interactions

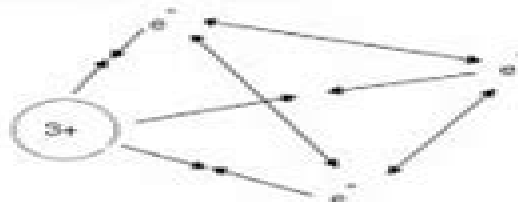
Reaction Equation for Electron Ionization: $X \rightarrow X^+ + e^-$ $\Delta E = 1^{\text{st}}$ ionization energy

Reaction Equation for Electron Capture: $X + e^- \rightarrow X^-$ $\Delta E = \text{electron affinity}$

Dominant inter-particle forces in helium:



Dominant inter-particle forces in lithium:



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Annelies Wilder-Smith



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, **Atomic Properties of the Elements**, 2001 *Understanding the Atom* U.S. Atomic Energy Commission, 1963
Chemistry and Our Universe, Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there

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