Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify the given (with units) and what you want to find (with units)
- c. Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

```
Equation: 2 \text{ H}_2\text{O}_{(1)} \rightarrow 2 \text{ H}_2(\omega) + 1 \text{ O}_2(\omega)

Before 2 \text{ mol} \quad 0 \text{ mol} \quad 0 \text{ mol}

Change -2 \text{ mol} \quad +2 \text{ mol}

After 0 \text{ mol} \quad 2 \text{ mol}
```

$$36 g H_2Ox \frac{1 mol H_2O}{18.0 g H_2O} = 2.0 mol H_2O \times \frac{2 mol H_2}{2 mol H_2O} = 2 mol H_2 \times \frac{2.0 g H_2}{1 mol H_2} = 4.0 g H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

```
      Equation:
      \frac{2}{2} Na<sub>(1)</sub> + \frac{1}{1} Cl<sub>2(2)</sub> \rightarrow \frac{2}{2} NaCl<sub>(n)</sub>

      Before
      9.94 mol
      xs mol
      0 mol

      Change
      -9.94 mol
      +9.94 mol

      After
      0 mol
      9.94 mol
```

$$581 \text{ g NaClx} \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

You eat 180.0 g of glucose (90 M&Ms). If glucose, C6H12O6, reacts with oxygen gas
to produce carbon dioxide and water, how many grams of oxygen will you have to

Percent Yield Stoichiometry Answers With Work

Chris Sunley

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