

7. Give a recipe for preparing a sodium carbonate/sodium bicarbonate buffer solution having a pH of 9.5. **For carbonic acid,  $\text{H}_2\text{CO}_3$ ,  $\text{pK}_{\text{a}1} = 6.35$  and  $\text{pK}_{\text{a}2} = 10.33$**

8. A saturated solution of silver chromate ( $\text{Ag}_2\text{CrO}_4$ , a sparingly soluble solid, formula mass 331.0 g/mol), was found to have a chromate ion concentration of  $3.2 \times 10^{-4} \text{ M}$  at  $25^\circ\text{C}$ . What is (a) the  $K_{\text{sp}}$ ; and (b) solubility of the salt in mg/mL? **How could you dissolve a sample of the salt?**

9. Calculate the solubility in mg/mL of  $\text{PbCl}_2$  ( formula mass 276.2 g/mol) in 0.25 M sodium chloride solution at  $25^\circ\text{C}$ .  $K_{\text{sp}}$  for lead chloride is  $1.8 \times 10^{-10}$  at  $25^\circ\text{C}$ .

10. Write an equation to show the formation of the complex ion  $[\text{Ag}(\text{NH}_3)_2]^+$  from silver ion and aqueous ammonia then calculate the concentration of silver ions in solution if exactly 0.100 mol of silver nitrate is added to exactly **500 mL of 0.800 M ammonia** solution.  $K_f$  for  $[\text{Ag}(\text{NH}_3)_2]^+$  is  $1.7 \times 10^7$ .

# Recipe For Sodium Carbonate Sodium Bicarbonate Buffer

**John Kiernan**



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